

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|----------------------------|-------------------|-------------------|
| acid dissociation constant | 1. conjugate base | 2. weak acid |
| acid ionization | 3. conjugate acid | 3. autoionization |
| acid ionization | 4. weak base | |

Key Equations

- | | |
|------------------------------------|-----------------------------------|
| 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ | 2. $K_b = \frac{[HB][OH^-]}{[B]}$ |
| 3. $K_a + K_b = K_w$ | |

Part A Completion

Write the appropriate words in the blanks to complete the following statements. Some blanks may require chemical formulas.

1. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
2. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
3. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
4. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
5. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
6. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
7. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
8. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

9. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
10. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
11. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
12. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
13. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
14. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .
15. The K_a of a weak acid is 1.0×10^{-4} . The K_b of its conjugate base is 1.0×10^{-10} .
16. The K_b of a weak base is 1.0×10^{-4} . The K_a of its conjugate acid is 1.0×10^{-10} .

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Daniela Niemeyer



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