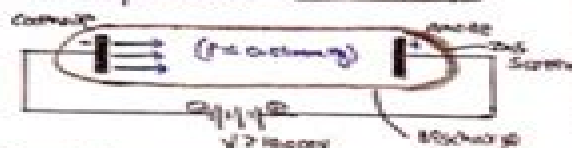


## Discovery of Fundamental Particles:

### Electron :-

- Discovered by J.J. Thomson in 'Cathode Ray Experiment'.
- When high voltage ( $>10^4$  V) is applied across a discharge tube containing a gas at very low pressure (a vacuum), then a ray flows from cathode to anode.

→ This ray is called 'Cathode Ray'.



### Properties of Cathode Rays :-

- 1) Travels in straight line with high speed (fast stream).
- 2) Deviates like particle in motion.
- 3) Carry charge (gets deflected towards the plate in electric field).
- 4) Carry momentum (deviates like particle).
- 5) Affects photographic plate and produces 'Green Glow' when strike wall beyond anode.
- 6) Can pass through metal sheets.
- 7) Composed of very charged electrons.

- Path of  $e^-$  in uniform EF is parabolic.
- Path of  $e^-$  in uniform mag. field is circular ( $r = \frac{mv}{eB}$ ).
- Charge by mass ratio ( $\frac{e}{m}$ ) =  $1.76 \times 10^{11}$  C/kg.
- Charge of  $e^-$  was discovered by Millikan oil exp.
- Finding mass of  $e^-$  ( $m_e$ ) =  $\frac{h}{\lambda} = \frac{h}{v\lambda}$  (from Planck's eqn).

- Characteristics of cathode rays are independent of gas nature of gas taken in discharge tube.

### Proton :-

- Discovered by Goldstein & Rutherford in 'Cathode Ray Experiment'.
- Experimental setup is same as above. The only diff. is that here instead of anode, cathode is perforated (cathode with holes).

→ 'Cathode/Anode Ray' flows from holes of Cathode.



## Atomic Structure - 2

### Properties of Cathode Rays :-

- 1) Travels in str. line with very high speed.
- 2) Carry +ve charge & momentum.
- 3) Behave like particle in motion.
- 4) Characteristics of anode rays depends upon nature of gas taken in discharge tube.
- 5) If the gas taken in discharge tube, then ray consists of +vely charged particles called 'proton'.

$$\text{Mass of proton} = 1.6726 \times 10^{-27} \text{ kg}$$

### Neutron :-

- Discovered by J. Chadwick.
- When Beryllium (Be) nucleus bombarded with  $\alpha$ -particles (He nuclei) then neutrons emitted.



Particle	Representation	Charge	Mass
$\alpha$ particle	${}^4_2\text{He}$	+2e	4 amu
$\beta$ particle	$e^-$	-e	$\frac{1}{1840}$ amu
Proton	${}^1_1\text{H}$ or $p^+$	+e	1.007 amu
Neutron	${}^1_0\text{n}$	zero	1.008 amu

Specific Charge =  $\frac{\text{Charge of } \alpha \text{ particle}}{\text{Mass of } \alpha \text{ particle}}$

Order of SC:  $e^- > p > \alpha \text{ particle} > n$

### Isotopes :-

- 1) Isotopes → Same Z but different A.  
Eq:  ${}^1_1\text{H}$ ,  ${}^2_1\text{H}$ ,  ${}^3_1\text{H}$
- 2) Isobars → Same A but different Z.  
Eq:  ${}^4_2\text{He}$ ,  ${}^4_1\text{H}$
- 3) Isodiaphers → Same no. of (n-p) or (A-2Z)  
Eq:  ${}^{14}_6\text{C}$ ,  ${}^{14}_7\text{N}$
- 4) Isotones → Same no. of neutrons (A-Z)  
Eq:  ${}^{14}_6\text{C}$ ,  ${}^{15}_7\text{N}$
- 5) Isotectonics Species → species having same number of  $e^-$   
Eq:  $\text{Na}^+$ ,  $\text{Mg}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{C}^{4+}$ ,  $\text{N}^{5+}$ ,  $\text{O}^{6+}$ ,  $\text{F}^{7+}$ ,  $\text{Ne}$
- 6) Isostere → Isoelectronic species having same no. of atoms/molecules.



### Thomson's Model of Atom :-

- Also known as plum pudding / watermelon model.
- According to J.J. Thomson, atom consists of uniform sphere of +ve charge in which the electrons are distributed uniformly.



### Drawbacks :-

- Mass considered evenly spread over the atom.
- It is a stable model. Doesn't explain movement of  $e^-$ .
- Can't explain stability of atom.
- Failed to explain Rutherford  $\alpha$ -particle experiment.

### Rutherford Model & $\alpha$ -particle Experiment :-

Rutherford bombarded  $\alpha$ -particle on a thin gold foil.



### Observations :-

- Most of the  $\alpha$ -particles pass through the gold foil undeviated.
- Few  $\alpha$ -particles get deflected by some angle.

→ Very few  $\alpha$ -particles (1 in 2000) returned back to its original path.

### Conclusions :-

- 1) Most of the atom is empty.
- 2) Major part by mass concentrated at center in very small region called 'nucleus'.
- 3) Nucleus is very small as compared to atom.

$$r_{\text{atom}} \approx 10^{-10} \text{ m}; \quad r_{\text{nucleus}} \approx 10^{-15} \text{ m}$$

- 4) Nucleus is +vely charged.
- 5) Protons and neutrons are present in nucleus.
- 6) Since an atom is neutral, in an atom No. of proton = No. of  $e^-$
- 7)  $e^-$  revolves around the nucleus in a definite path called orbits.

### Distance of Closest Approach :-

→ Minimum distance from nucleus at which  $\alpha$ -particle stops due to repulsion & goes back.



$$R = \frac{kZe^2}{mv^2}$$

# Advanced Physical Chemistry Notes

**BM King**



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