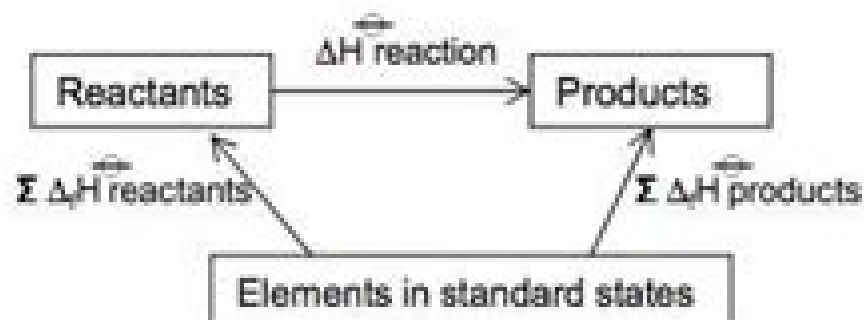


# Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$



**Example 6.** What is the enthalpy change for this reaction ?



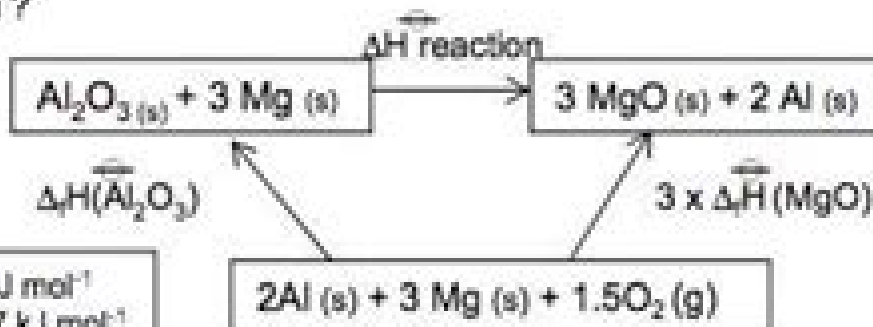
$$\Delta H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta_f H (\text{MgO}) - \Delta_f H (\text{Al}_2\text{O}_3)$$

$$\begin{aligned} \Delta H &= (3 \times -601.7) - -1675.7 \\ &= -129.4 \text{ kJ mol}^{-1} \end{aligned}$$

Remember  
elements have  
 $\Delta_f H = 0$

$$\begin{aligned} \Delta_f H (\text{MgO}) &= -601.7 \text{ kJ mol}^{-1} \\ \Delta_f H (\text{Al}_2\text{O}_3) &= -1675.7 \text{ kJ mol}^{-1} \end{aligned}$$



**Example 7.** Using the following data to calculate the heat of combustion of propene

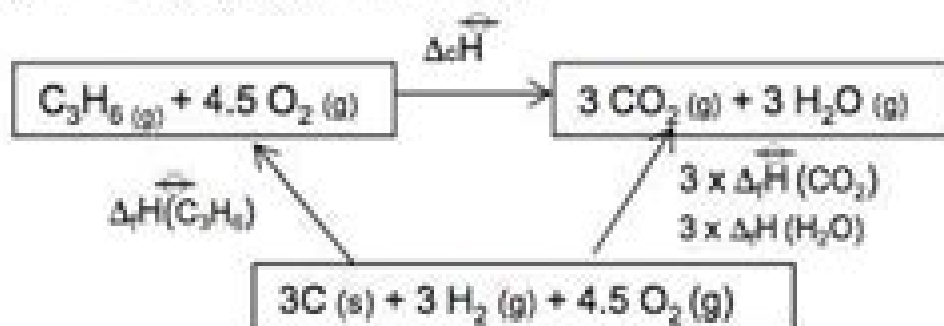
$$\Delta_f H, \text{C}_3\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$



$$\Delta_c H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta_c H = [3 \times \Delta_f H (\text{CO}_2) + 3 \times \Delta_f H (\text{H}_2\text{O})] - \Delta_f H (\text{C}_3\text{H}_6)$$

$$\begin{aligned} \Delta_c H &= [(3 \times -394) + (3 \times -242)] - 20 \\ &= -1928 \text{ kJ mol}^{-1} \end{aligned}$$



# Answers For Chemfax Thermodynamics Enthalpy Lab

**Marc J. Assael, William A.  
Wakeham, Anthony R. H.  
Goodwin, Stefan Will, Michael  
Stamatoudis**

## Answers For Chemfax Thermodynamics Enthalpy Lab:

**Paperbound Books in Print 1995** Reed Reference Publishing, R5ference Reed, 1995-12      Commonly Asked Questions in Thermodynamics Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-10

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undergraduate course or not Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems **Hess's Law** Lifeliqe,2019 This lesson plan covers two indirect methods of determining enthalpy change for most chemical reactions using Hess s law

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