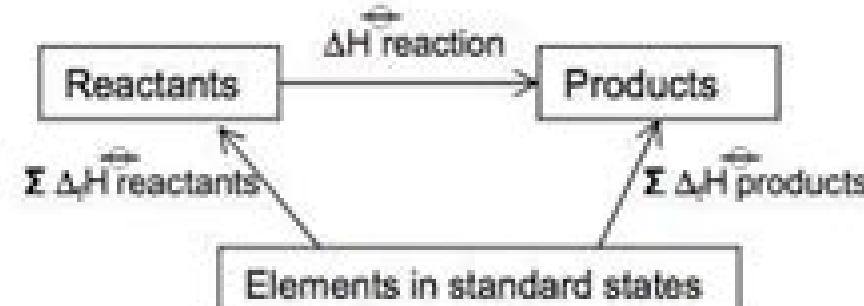


Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$$



Example 6. What is the enthalpy change for this reaction ?



$$\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta H(\text{MgO}) - \Delta H(\text{Al}_2\text{O}_3)$$

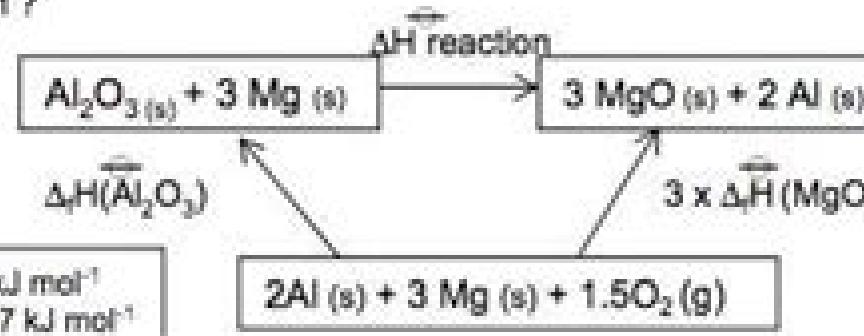
$$\Delta H = (3 \times -601.7) - -1675.7$$

$$= -129.4 \text{ kJ mol}^{-1}$$

Remember elements have $\Delta H_f = 0$

$$\Delta H(\text{MgO}) = -601.7 \text{ kJ mol}^{-1}$$

$$\Delta H(\text{Al}_2\text{O}_3) = -1675.7 \text{ kJ mol}^{-1}$$



Example 7. Using the following data to calculate the heat of combustion of propene

$$\Delta H_f, \text{C}_2\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta H_f, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta H_f, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$

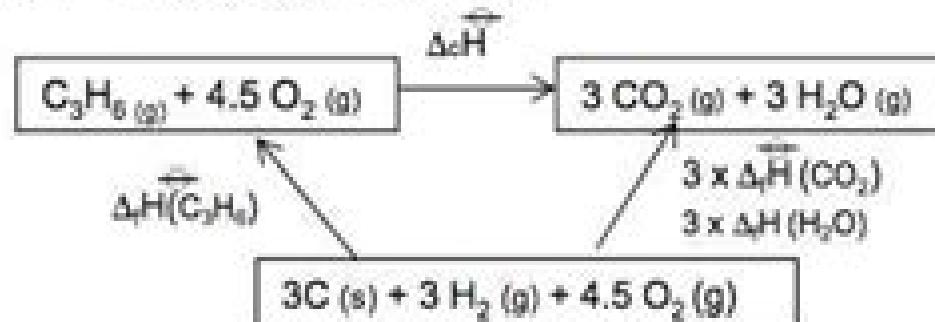


$$\Delta H = \sum \Delta H_{\text{products}} - \sum \Delta H_{\text{reactants}}$$

$$\Delta H = [3 \times \Delta H(\text{CO}_2) + 3 \times \Delta H(\text{H}_2\text{O})] - \Delta H(\text{C}_2\text{H}_6)$$

$$\Delta H = [(3 \times -394) + (3 \times -242)] - 20$$

$$= -1928 \text{ kJ mol}^{-1}$$



Answers For Chemfax Thermodynamics Enthalpy Lab

**Marc J. Assael, William A.
Wakeham, Anthony R. H.
Goodwin, Stefan Will, Michael
Stamatoudis**

Answers For Chemfax Thermodynamics Enthalpy Lab:

Paperbound Books in Print 1995 Reed Reference Publishing, Reference Reed, 1995-12 [Commonly Asked Questions in Thermodynamics](#) Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-10

Have you ever had a question that keeps persisting and for which you cannot find a clear answer? Is the question seemingly so simple that the problem is glossed over in most resources or skipped entirely? CRC Press Taylor and Francis is pleased to introduce *Commonly Asked Questions in Thermodynamics*, the first in a new series of books that address the questions that frequently arise in today's major scientific and technical disciplines. Designed for a wide audience from students and researchers to practicing professionals in related areas, the books are organized in a user friendly Question Answer format. Presented questions become increasingly specific throughout the book with clear and concise answers as well as illustrations. Diagrams and tables are incorporated wherever helpful. Thermodynamics is a core discipline associated with the theoretical principles and practical applications underlying almost every area of science from nanoscale biochemical engineering to astrophysics. Highlighting chemical thermodynamics in particular, this book is written in an easy to understand style and provides a wealth of fundamental information, simple illustrations and extensive references for further research and collection of specific data. Designed for an audience that ranges from undergraduate students to scientists and engineers at the forefront of research, this indispensable guide presents clear explanations for topics with wide applicability. It reflects the fact that very often the most common questions are also the most profound.

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undergraduate course or not. Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems. **Hess's Law** Lifeliqe, 2019. This lesson plan covers two indirect methods of determining enthalpy change for most chemical reactions using Hess's law.

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