

Chapter 1 - Chemical Foundations

1. Consider the element indium, atomic number 49, atomic mass 114.8 g. The nucleus of an atom of indium-112 contains:

- a. 49 protons, 63 neutrons, 49 electrons.
- b. 49 protons, 49 neutrons.
- c. 49 protons, 49 alpha particles.
- d. 49 protons, 63 neutrons.
- e. 49 protons, 112 neutrons.

ANSWER:

d

POINTS:

1

2. A hypothetical element consists of two isotopes of masses 78.95 amu and 80.95 amu with abundances of 32.9% and 67.1%, respectively. What is the average atomic mass of this element?

- a. 79.95 amu
- b. 78.95 amu
- c. 79.6 amu
- d. 80.3 amu
- e. 80.95 amu

ANSWER:

d

POINTS:

1

3. Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29 (28.976 amu, 4.67% abundance), and X-30 (29.974 amu, 3.10% abundance). Calculate the atomic weight of X.

- a. 28.09 amu
- b. 48.63 amu
- c. 27.16 amu
- d. 28.97 amu
- e. 86.93 amu

ANSWER:

a

POINTS:

1

4. The atomic mass of rhenium is 186.2. Given that 37.1% of natural rhenium is rhenium-185, what is the other stable isotope?

- a. $^{183}_{75}\text{Re}$
- b. $^{187}_{75}\text{Re}$
- c. $^{189}_{75}\text{Re}$
- d. $^{181}_{75}\text{Re}$
- e. $^{180}_{75}\text{Re}$

ANSWER:

b

POINTS:

1

5. Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29

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Kenneth W. Whitten

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