

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

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## **Table of Contents Answers Chapter Review Test Holt The Atmosphere**

1. Understanding the eBook Answers Chapter Review Test Holt The Atmosphere
  - The Rise of Digital Reading Answers Chapter Review Test Holt The Atmosphere
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers Chapter Review Test Holt The Atmosphere
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers Chapter Review Test Holt The Atmosphere
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Chapter Review Test Holt The Atmosphere
  - Personalized Recommendations

- Answers Chapter Review Test Holt The Atmosphere User Reviews and Ratings
- Answers Chapter Review Test Holt The Atmosphere and Bestseller Lists
- 5. Accessing Answers Chapter Review Test Holt The Atmosphere Free and Paid eBooks
  - Answers Chapter Review Test Holt The Atmosphere Public Domain eBooks
  - Answers Chapter Review Test Holt The Atmosphere eBook Subscription Services
  - Answers Chapter Review Test Holt The Atmosphere Budget-Friendly Options
- 6. Navigating Answers Chapter Review Test Holt The Atmosphere eBook Formats
  - ePub, PDF, MOBI, and More
  - Answers Chapter Review Test Holt The Atmosphere Compatibility with Devices
  - Answers Chapter Review Test Holt The Atmosphere Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers Chapter Review Test Holt The Atmosphere
  - Highlighting and Note-Taking Answers Chapter Review Test Holt The Atmosphere
  - Interactive Elements Answers Chapter Review Test Holt The Atmosphere
- 8. Staying Engaged with Answers Chapter Review Test Holt The Atmosphere
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers Chapter Review Test Holt The Atmosphere
- 9. Balancing eBooks and Physical Books Answers Chapter Review Test Holt The Atmosphere
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers Chapter Review Test Holt The Atmosphere
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Chapter Review Test Holt The Atmosphere
  - Setting Reading Goals Answers Chapter Review Test Holt The Atmosphere
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Chapter Review Test Holt The Atmosphere
  - Fact-Checking eBook Content of Answers Chapter Review Test Holt The Atmosphere

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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