

STOICHIOMETRY & LIMITING REAGENT

- Q.1 If 0.5 moles of BaCl_2 is mixed with 0.2 moles of Na_3PO_4 , the maximum number of moles of $\text{Ba}_3(\text{PO}_4)_2$ that can be formed is - (A) 0.2 (B) 0.5 (C) 0.3 (D) 0.1 [D]
- Q.2 8.4 g MgCO_3 on heating leaves behind a residue weighing 4.0 g carbon dioxide released into the atmosphere at S.T.P. will be - (A) 2.24 L (B) 4.48 L (C) 1.12 L (D) 0.56 L [A]
- Q.3 5 moles of A, 6 moles of Z and mixed with sufficient amount of C to produce final product F. How many maximum moles of 'F' can be produced as per the given sequence of reaction ?
 $\text{A} + 2\text{Z} \rightarrow \text{B}$;
 $\text{B} + \text{C} \rightarrow \text{Z} + \text{F}$ (A) 3 (B) 2 (C) 4 (D) 5 [A]
- Q.4 The chloride of a metal contains 71% chlorine by weight and the vapour density of it is 50. The atomic weight of the metal will be (valency of metal is 2) - (A) 29 (B) 58 (C) 35.5 (D) 71 [A]
- Q.5 A reaction required three atoms of Mg for two atoms of N. How many gm of N are required for 3.6 gm of Mg ? (A) 2.43 (B) 4.86 (C) 1.4 (D) 4.25 [C]
- Q.6 510 mg of liquid on vaporization in Victor Meyer's apparatus displaces 67.2 cm^3 of air at (STP). The molecular weight of the liquid is - (A) 130 (B) 17 (C) 170 (D) 1700 [C]
- Q.7 When 3.2 g S is vapourized at 450°C and 723 mm pressure, the vapours occupy a volume of 780 ml. What is the molecular formula of S vapours under these conditions
 (A) S_2 (B) S_4 (C) S_6 (D) S_8 [D]
- Q.8 2.76 g of silver carbonate on being strongly heated yields a residue weighing -
 (A) 2.16 g (B) 2.48 g (C) 2.32 g (D) 2.64 g [A]
- Q.9 3 gm of Mg is burnt in a closed vessel containing 3 gm of oxygen. The weight of excess reactant left is-
 (A) 0.5 gm of oxygen (B) 1.0 gm of oxygen (C) 1.0 gm of Mg (D) 0.5 gm of Mg [B]
- Q.10 0.54 gm of metal "M" yields 1.02 gm of its oxide M_2O_3 . The at. wt. of metal "M" is-
 (A) 9 (B) 18 (C) 27 (D) 54 [C]
- Q.11 The mass of oxygen that would be required to produce enough CO, which completely reduces 1.6 kg Fe_2O_3 (at. mass Fe = 56) is- (A) 240 gm (B) 480 gm (C) 720 gm (D) 960 gm [B]
- Q.12 If one mole of ethanol ($\text{C}_2\text{H}_5\text{OH}$) completely burns to carbon dioxide and water, the weight of carbon dioxide formed is about - (A) 22g (B) 45g (C) 66g (D) 88g [D]
- Q.13 12 litre of H_2 and 11.2 litre of Cl_2 are mixed and exploded. The composition by volume of mixture is-
 (A) 24 litre of HCl (B) 0.8 litre Cl_2 and 20.8 lit HCl
 (C) 0.8 litre H_2 & 22.4 litre HCl (D) 22.4 litre HCl [C]
- Q.14 The number of atoms of Cr and O are 4.8×10^{10} and 9.6×10^{10} respectively. Its empirical formula is -
 (A) Cr_2O_3 (B) CrO_2 (C) Cr_2O_4 (D) none [B]
- Q.15 16 gm of SO_x occupies 5.6 litre at STP. Assuming ideal gas nature, the value of x is -
 (A) 1 (B) 2 (C) 3 (D) None [B]
- Q.16 Number of moles of water in 488 gm of $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ are - (Ba = 137)
 (A) 2 moles (B) 4 moles (C) 3 moles (D) 5 moles [B]
- Q.17 14g of element X combine with 16g of oxygen. On the basis of this information, which of the following is a correct statement :
 (A) The element X could have an atomic weight of 7 and its oxide formula XO
 (B) The element X could have an atomic weight of 14 and its oxide the formula X_2O

Answer Sheet To The Stoichiometry Limiting Reagent

John T. Moore, Richard H. Langley



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