

Experiment 9

Advance Study Assignment: Molar Mass of a Volatile Liquid

1. A student weighed an empty flask and stopper and found the mass to be 53.256 g. She then added about five milliliters of an unknown liquid and heated the flask (in a boiling-water bath) to 98.8°C. After all the liquid had vaporized, she removed the flask from the bath, stoppered it, and let it cool. Once it had cooled, she momentarily removed the stopper, then replaced it and weighed the flask and condensed vapor, obtaining a mass of 53.870 g. The volume of the flask was known to be 271.1 mL. The absolute atmospheric pressure in the laboratory on that day was 728 mm Hg.

a. What was the pressure of the vapor in the flask in atm?

$$P = \underline{\hspace{2cm}} \text{ atm}$$

b. What was the temperature of the vapor in K? What was the volume of the flask in liters?

$$T = \underline{\hspace{2cm}} \text{ K} \quad V = \underline{\hspace{2cm}} \text{ L}$$

c. What was the mass of condensed vapor that was present in the flask?

$$m = \underline{\hspace{2cm}} \text{ grams}$$

d. How many moles of condensed vapor were present?

$$n = \underline{\hspace{2cm}} \text{ moles}$$

Advance Study Assignment Experiment Molar Mass

M Lipman

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