

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

Rhe = A

$$\lambda = 2.18 \times 10^{-8} \text{ m}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

$$h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$$

$$c = 3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?
 a) 0.307 m d) 0.110 m
 b) 0.0365 m $\frac{3.00 \times 10^8 \text{ m}}{8.22 \times 10^9 \text{ s}^{-1}} = 27.4 \text{ m}$
 c) 0.122 m $\frac{3.00 \times 10^8 \text{ m}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.36496 \text{ m}$

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?
 a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$
 b) 3.30 m d) 2.73 m
 $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$

3. Which one of the following is NOT a proper unit for frequency?
 a) Hz c) $\text{m} \cdot \text{s}^{-1}$ *velocity*
 b) s^{-1} d) $\frac{1}{\text{sec}}$

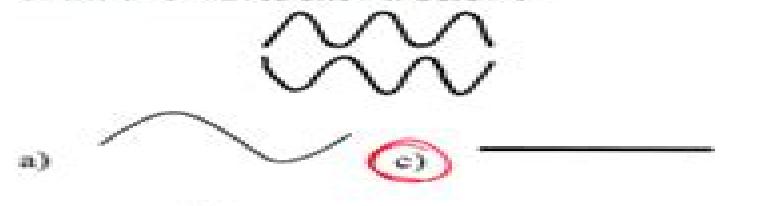
4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.
 $E_4 = -2.18 \times 10^{-18} \text{ J}$ $E_2 = -2.18 \times 10^{-13} \text{ J}$
 $\Delta E = 4.84 \times 10^{-19} \text{ J}$ $\lambda = \frac{hc}{\Delta E} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{4.84 \times 10^{-19} \text{ J}} = 4.1029 \times 10^{-7} \text{ m}$

5. What is the relationship between the energy of a photon of light and its frequency?
 a) $E = \nu$ d) $E = \frac{1}{\nu}$
 b) $E = \frac{h}{\nu}$ c) $E = \frac{\nu}{h}$
 e) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?
 a) $1.52 \times 10^{-4} \text{ J}$ d) $4.48 \times 10^{-19} \text{ J}$
 b) $3.63 \times 10^{-19} \text{ J}$ e) $3.03 \times 10^{-19} \text{ J}$
 $\Delta E = E_3 - E_2 = \frac{hc}{\lambda_3} - \frac{hc}{\lambda_2} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{9.11 \times 10^{-8} \text{ m}} - \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{1.097 \times 10^7 \text{ m}^{-1}}$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light?
 a) $3.03 \times 10^{-12} \text{ m}$ c) $3.30 \times 10^{11} \text{ m}$
 b) $2.42 \times 10^{-12} \text{ m}$ d) $1.59 \times 10^{-25} \text{ m}$
 $\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{7.14 \times 10^5 \text{ kg} \cdot 8.0 \times 3.00 \times 10^8 \text{ m/s}}$

8. What resultant is expected from the interference of the two waves shown below?



9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?
 a) α c) m_f
 b) ℓ d) m_s

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