

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

Rhe = A

$$\lambda = 2.18 \times 10^{-8} \text{ m}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

$$h = 6.626 \times 10^{-34} \text{ J} \cdot \text{s}$$

$$c = 3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?
 a) 0.307 m d) 0.110 m
 b) 0.0365 m $\frac{3.00 \times 10^8 \text{ m}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$
 c) 0.122 m

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?
 a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$
 b) 3.30 m d) 2.73 m
 $\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}}{110 \times 10^6 \text{ Hz}} = 2.73 \text{ m}$

3. Which one of the following is NOT a proper unit for frequency?
 a) Hz c) $\text{m} \cdot \text{s}^{-1}$ *velocity*
 b) s^{-1} d) $\frac{1}{\text{sec}}$

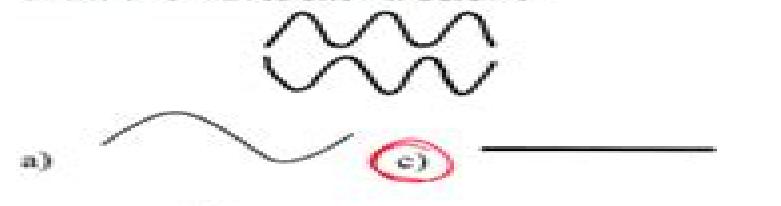
4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.
 $E_4 = -2.18 \times 10^{-13} \text{ J}$ $E_2 = -2.18 \times 10^{-13} \text{ J}$
 $\Delta E = 4.84 \times 10^{-19} \text{ J}$ $\lambda = \frac{hc}{\Delta E} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{4.84 \times 10^{-19} \text{ J}} = 4.1029 \times 10^{-7} \text{ m}$

5. What is the relationship between the energy of a photon of light and its frequency?
 a) $E = \nu$ d) $E = \frac{1}{\nu}$
 b) $E = \frac{h}{\nu}$ c) $E = \frac{\nu}{h}$
 e) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?
 a) $1.52 \times 10^4 \text{ J}$ d) $4.48 \times 10^{-19} \text{ J}$
 b) $3.63 \times 10^{-19} \text{ J}$ e) $3.03 \times 10^{-19} \text{ J}$
 $\Delta E = E_3 - E_2 = \frac{hc}{\lambda_3} - \frac{hc}{\lambda_2} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{7.14 \times 10^{-15} \text{ m}} - \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s} \cdot 3.00 \times 10^8 \text{ m}}{4.84 \times 10^{-19} \text{ m}}$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light?
 a) $3.03 \times 10^{-12} \text{ m}$ c) $3.30 \times 10^{11} \text{ m}$
 b) $2.42 \times 10^{-12} \text{ m}$ d) $1.59 \times 10^{-25} \text{ m}$
 $\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{7.14 \times 10^{-15} \text{ kg} \cdot 8.0 \times 3.00 \times 10^8 \text{ m/s}}$

8. What resultant is expected from the interference of the two waves shown below?



9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?
 a) m_s c) m_l
 b) ℓ d) m_s

Ap Chemistry Worksheet Answers

Oscar Krisen Buros

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Carolina Science and Math Carolina Biological Supply Company,2003 *Principles of Physical Chemistry* Hans Kuhn,David H. Waldeck,Horst-Dieter Försterling,2024-10-25 Core textbook showcasing the broad scope and coherence of physical chemistry Principles of Physical Chemistry introduces undergraduate students to the concepts and methods of physical chemistry which are fundamental to all of Chemistry In their unique approach the authors guide students along a logically consistent pathway from the principles of quantum mechanics and molecular structure to the properties of ensembles and supramolecular machines with many examples from biology and nanoscience By systematically proceeding from atoms to increasingly complex forms of matter the book elucidates the connection between recognizable paradigms and modern chemistry research in a student friendly manner To promote intuition and understanding for beginning students the text introduces concepts before proceeding to more rigorous treatments Rigorous proofs and derivations are provided as electronic supplements for more advanced students The book poses over 900 exercises and problems to help the student learn and master methods for physicochemical reasoning Computational supplementary material including Fortran simulations MathCAD exercises and Mathematica programs are included on a companion website Some topics discussed in the text are Electronic structure and Variational Principle including Pauli exclusion spin orbit interactions and electron confinement in quantum dots Chemical bonding and molecular structure including electron tunneling comparison of electron in a box models and electron orbital methods and the mechanics of chemical bonds Absorption and emission of light including transition dipoles for electron systems coupled chromophores excitons and chiroptical activity Statistical description of molecular ensembles including microscopic interpretations of phase transitions entropy work and heat Chemical equilibria including statistical description of equilibrium constants electrochemistry and the exposition of fundamental reaction types Reaction kinetics and reaction dynamics including nonlinear coupled reactions femtochemistry and solvent effects on reactions Physicochemical properties of macromolecules and the principles of supramolecular assemblies including polymer dynamics and chemical control of interfaces The logic of supramolecular machines and their manipulation of photon electron and nuclear motion With its highly coherent and systematic approach to the subject Principles of Physical Chemistry is an ideal textbook and resource for students in undergraduate physical chemistry courses especially those in programs of study related to chemistry engineering and molecular and chemical biology **The Art of Teaching Science** Jack Hassard,2005 This is a core teaching textbook designed for the professional development of middle and high school science teachers Differing from other texts in its constructivist approach to learning and teaching it provides meaningful learning experiences and connections with the most recent research and understanding of science teaching Each chapter is organized into two sections the first focuses on the content of the major theme of the chapter while the second consists of a newspaper like feature called The Science Teaching Gazette containing a variety of strategies for extending the

learning process Packed with learning tools hands on inquiry activities case studies think pieces and interviews with teachers around the world this is a remarkably comprehensive textbook designed to prepare a new cadre of science teachers Midwest

Teaching Better Bradley A. Ermeling, Genevieve Graff-Ermeling, 2016-03-03 Discover the power of collaborative inquiry This unique visually stunning resource is packed with details to ignite and sustain the collaborative improvement of teaching and learning Includes US and international case studies powerful metaphors application exercises a leader's guide a companion website digital templates and more Learn what lesson study and collaborative inquiry can and should look like Find the guidance you need to lead and support schoolwide inquiry based improvement A true inspiration for educators who want to improve both their own craft and the methods of the profession Jim Stigler James Hiebert Authors of The Teaching Gap **Illinois Chemistry Teacher**, 2007-09 Journal of Biological Education, 1989 Defining Student Success Lisa

M. Nunn, 2014-04-15 The key to success our culture tells us is a combination of talent and hard work Why then do high schools that supposedly subscribe to this view send students to college at such dramatically different rates Why do students from one school succeed while students from another struggle To the usual answer an imbalance in resources this book adds a far more subtle and complicated explanation Defining Student Success shows how different schools foster dissimilar and sometimes conflicting ideas about what it takes to succeed ideas that do more to preserve the status quo than to promote upward mobility Lisa Nunn's study of three public high schools reveals how students' beliefs about their own success are shaped by their particular school environment and reinforced by curriculum and teaching practices While American culture broadly defines success as a product of hard work or talent at school intelligence is the talent that matters most Nunn shows that each school refines and adapts this American cultural wisdom in its own distinct way reflecting the sensibilities and concerns of the people who inhabit each school While one school fosters the belief that effort is all it takes to succeed another fosters the belief that hard work will only get you so far because you have to be smart enough to master course concepts Ultimately Nunn argues that these school level adaptations of cultural ideas about success become invisible advantages and disadvantages for students' college going futures Some schools' definitions of success match seamlessly with elite college admissions' definition of the ideal college applicant while others more closely align with the expectations of middle or low tier institutions of higher education With its insights into the transmission of ideas of success from society to school to student this provocative work should prompt a reevaluation of the culture of secondary education Only with a thorough understanding of this process will we ever find more consistent means of inculcating success by any measure

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