

AP CHEMISTRY CHAPTER 6 PRACTICE TEST

- For which chemical equation will the heat of reaction ΔH° , correspond to the molar heat of formation of liquid water?
 - $H_{2(g)} + 1/2O_{2(g)} \rightarrow H_2O_{(l)}$
 - $H_{2(g)} + H_{2(g)} + O_{2(g)} \rightarrow H_2O_{(l)}$
 - $2H_{2(g)} + O_{2(g)} \rightarrow 2H_2O_{(l)}$
 - $H_{(aq)} + OH_{(aq)} \rightarrow H_2O_{(l)}$
- For the process $H_{2(g)} + 1/2O_{2(g)} \rightarrow H_2O_{(l)}$, $\Delta H = -247 \text{ kJ}$. The numeric value of -242 kJ represents heat energy that is
 - absorbed by the reaction
 - required to break the H-H and O=O bond
 - stored as potential energy
 - released by the reaction
- An 11.0 g sample of propane is placed in a bomb calorimeter with excess oxygen and is ignited. The initial temperature of the calorimeter is 25.00°C and its total heat capacity is 97.1 kJ/°C. The reaction raises the temperature of the calorimeter is 30.75°C. The approximate molar heat of combustion for propane is
 - 50.6 kJ/mol
 - 558 kJ/mol
 - $-2.24 \times 10^3 \text{ kJ/mol}$
 - $-1.20 \times 10^4 \text{ kJ/mol}$
- What is the change in enthalpy at 25°C and 1 atm for the production of 3.00 mol $SnO_{(s)}$?
$$Sn_{(s)} + SnO_{2(s)} \rightarrow 2SnO_{(s)} \quad \Delta H^\circ = 16.2 \text{ kJ}$$
 - 24.3 kJ
 - 16.2 kJ
 - 5.40 kJ
 - 16.2 kJ
 - 24.3 kJ
- How much heat is liberated at constant pressure if 0.896 g of calcium carbonate reacts with 92.3 mL of 0.584 M hydrochloric acid?
$$CaCO_{3(s)} + 2HCl_{(aq)} \rightarrow CaCl_{2(aq)} + H_2O_{(l)} + CO_{2(g)} \quad \Delta H^\circ = -15.2 \text{ kJ}$$
 - 0.410 kJ
 - 8.88 kJ
 - 13.6 kJ
 - 0.546 kJ
 - 0.136 kJ

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