

# AP CHEMISTRY CHAPTER 6 PRACTICE TEST

- For which chemical equation will the heat of reaction  $\Delta H^\circ$ , correspond to the molar heat of formation of liquid water?
  - $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $\text{H}_{2(g)} + \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$
  - $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
  - $\text{H}^+_{(aq)} + \text{OH}^-_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)}$
- For the process  $\text{H}_{2(g)} + 1/2\text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(l)}$ ,  $\Delta H = -247 \text{ kJ}$ . The numeric value of  $-242 \text{ kJ}$  represents heat energy that is
  - absorbed by the reaction
  - required to break the H-H and O=O bond
  - stored as potential energy
  - released by the reaction
- An 11.0 g sample of propane is placed in a bomb calorimeter with excess oxygen and is ignited. The initial temperature of the calorimeter is  $25.00^\circ\text{C}$  and its total heat capacity is  $97.1 \text{ kJ}/^\circ\text{C}$ . The reaction raises the temperature of the calorimeter to  $30.75^\circ\text{C}$ . The approximate molar heat of combustion for propane is
  - $-50.6 \text{ kJ/mol}$
  - $-558 \text{ kJ/mol}$
  - $-2.24 \times 10^3 \text{ kJ/mol}$
  - $-1.20 \times 10^4 \text{ kJ/mol}$
- What is the change in enthalpy at  $25^\circ\text{C}$  and 1 atm for the production of 3.00 mol  $\text{SnO}_{2(s)}$ ?
 
$$\text{Sn}_{(s)} + \text{SnO}_{2(s)} \rightarrow 2\text{SnO}_{(s)} \quad \Delta H^\circ = 16.2 \text{ kJ}$$
  - $24.3 \text{ kJ}$
  - $16.2 \text{ kJ}$
  - $5.40 \text{ kJ}$
  - $-16.2 \text{ kJ}$
  - $-24.3 \text{ kJ}$
- How much heat is liberated at constant pressure if 0.896 g of calcium carbonate reacts with 92.3 mL of 0.584 M hydrochloric acid?
 
$$\text{CaCO}_{3(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{CaCl}_{2(aq)} + \text{H}_2\text{O}_{(l)} + \text{CO}_{2(g)} \quad \Delta H^\circ = -15.2 \text{ kJ}$$
  - $-0.410 \text{ kJ}$
  - $-8.88 \text{ kJ}$
  - $-13.6 \text{ kJ}$
  - $-0.546 \text{ kJ}$
  - $-0.136 \text{ kJ}$

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