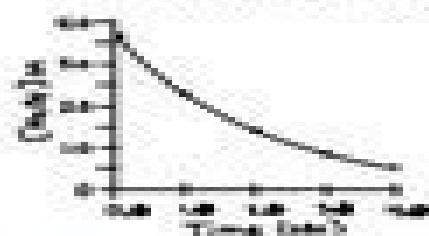
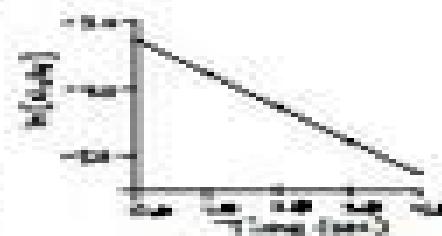


Q. For each of the graphs given below, determine the following:

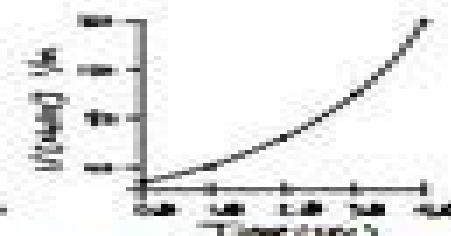
- The order of reaction (assume the other parameters except time are constant)
- The overall rate law
- The rate constant (to three decimal places)



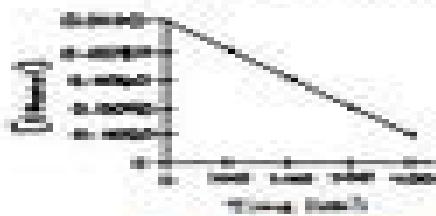
1st order



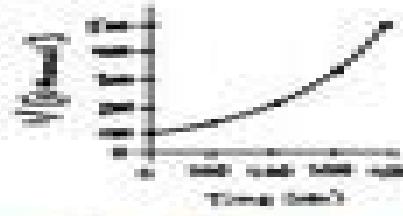
$$R = k[N_2O]$$



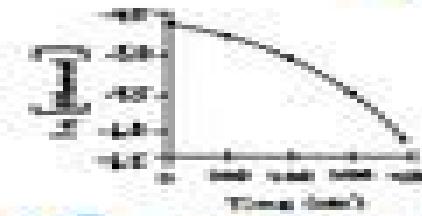
$$k = \text{slope} = \left(\frac{-0.3 - (-0.5)}{400 - 0} \right) = 0.53 \frac{1}{\text{sec}}$$



1st order



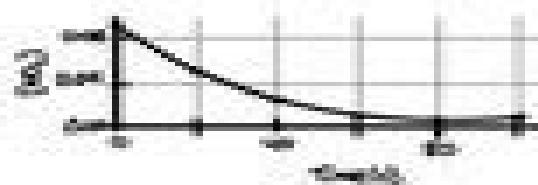
$$R = k[\text{CH}_3\text{CO}_2\text{H}]^2 \quad \text{or} \quad R = k$$



$$k = \text{slope} = \left(\frac{0.25 - 0.50}{400 - 0} \right) = 2 \times 10^{-5} \frac{1}{\text{sec}}$$



1st order



$$k = \text{slope} = \left(\frac{0.25 - 100}{400 - 40} \right) \approx 1.67 \frac{1}{\text{sec}}$$

Ans 2 min

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