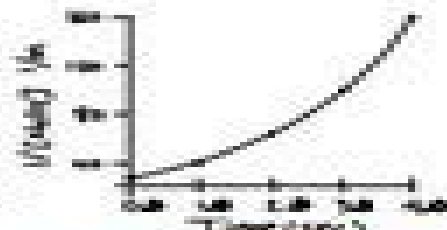
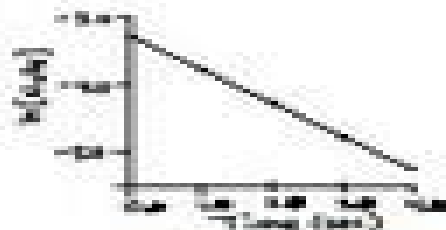
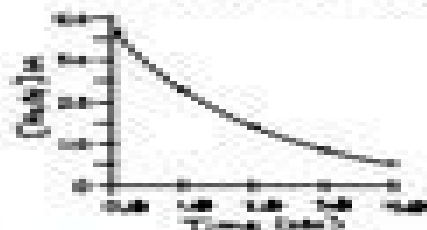


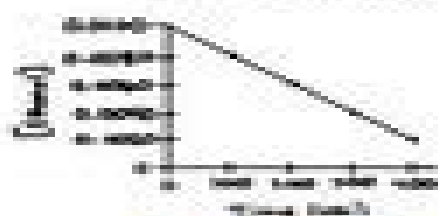
5. For each of the graphs set out below, determine the following:
 The order of the reaction (assume no other reactants impact the overall rate)
 The overall rate law
 The rate constant k with units



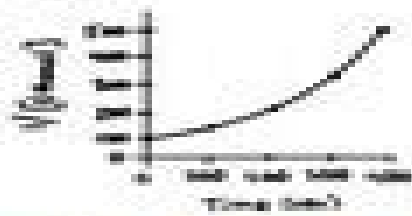
1st order

$$R = k[\text{N}_2\text{O}_5]$$

$$k = \text{slope} = \left(\frac{-3.3 - (-2.8)}{400 - 0} \right) = \boxed{0.53 \frac{1}{\text{sec}}}$$

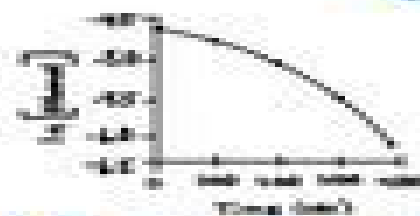


Zero order

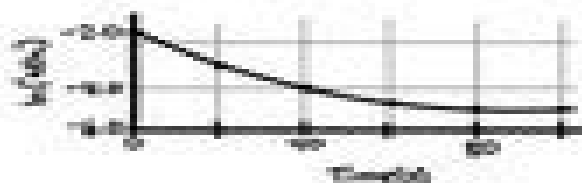


$$R = k[\text{CH}_3\text{CHO}]^0$$

or $R = k$

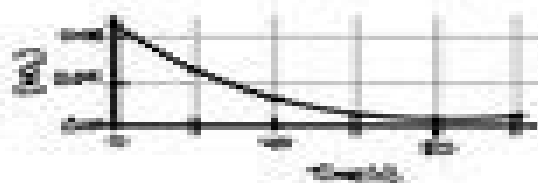


$$k = \text{slope} = \left(\frac{0.001 - 0.002}{1000 - 0} \right) = \boxed{2 \times 10^{-6} \frac{\text{mol}}{\text{L} \cdot \text{sec}}}$$



2nd order

$$R = k[\text{NO}_2]^2$$



$$k = \text{slope} = \left(\frac{100 - 10}{200 - 0} \right) = \boxed{1.67 \frac{1}{\text{mol} \cdot \text{s}}}$$

~~1st~~ \rightarrow 2 $\frac{1}{\text{mol}}$

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