

Chapter 2 Review

1. Which properties are characteristics of metals?
Which property means that you can punch the substance into a ball?

a) ductility b) malleability
c) conductivity d) brittleness

2. Which of the following is a transition metal?
a) Cu b) H c) Pb d) He

3. Which of the following is a transition metal?
a) Cu b) P c) Cu d) U

4. Which of the following are alkali metals?
a) Mg b) Na c) Si d) Al e) Li

5. Which of the following is an halogen?
a) Se b) Br c) Cl d) F e) W

6. Which element has the highest melting point?
a) Pb b) Au c) Cu d) Mg
e) Mg f) Au g) Cu h) Pb

7. Which term is not the
negative electrode? a) positive electrode
b) power source d) gas inside the tube

8. In a cathode ray tube electrons are held toward
a positively charged plate
b) a negatively charged plate

9. Listed below are the charges and masses of four
particles. Which one will be deflected by a magnetic
spectrometer?
a) +2.7 mass
b) -2.9 mass

c) +1.5 mass

Name _____
Period _____ Date _____

10. In a addition of long type expression, the charge
on the left of double Coulomb's law has to be

1.00000	Coulomb	0.00 Coulomb	1
1.00000	Coulomb	0.00 Coulomb	2
1.00000	Coulomb	0.00 Coulomb	3
1.00000	Coulomb	0.00 Coulomb	4

What is the charge of the electron according to this
expression?
a) 0.11 Coulomb b) 0.00 Coulomb
c) 0.07 Coulomb d) 0.00000 Coulomb

11. Picture below is a schematic of the Rutherford
experiment. Which numbered particle gives the
best evidence for the nuclear theory?



12. Which of the following is an isotope of the element
with 20 protons (20 protons = 20 neutrons = 20)?
a) Element 20 b) Element 20
c) isotope 20 d) Element 40

13. The radioactive element X has the following natural
abundance and average masses. What is the mass
of X?
a) 1% 20.07 mass 20.0%
b) 1% 20.09 mass 20.0%

Show your work.

$$(4.00)(0.00) + (20.10)(0.00) =$$

$$25.3_{\text{percent}}$$

For questions 14 - 17, use the following key:
which answer may be used once, more than once,
or not at all

a) alpha
b) beta
c) gamma
d) alpha and beta, but not gamma

14. a high energy beam of light
15. free electrons & free protons
16. a high speed electron
17. used by Ernest Rutherford as a "probe"

For questions 18 - 21, use the following key:
which answer may be used once, more than once,
or not at all

a) John Dalton
b) Ernest Rutherford
c) J.J. Thomson
d) Democritus

18. His model of the atom has been called the "plum
pudding" model

19. His model of the atom has been called the "billiard
ball" model

20. His model name is called the "atom".

21. His philosophical idea included the term "atom".

22. He added to the atomic theory the idea that atoms
had positive and negative parts.

23. Consider the following sentence. In the
blank, indicate below is present:
a) the particle contains 30 protons

b) The particle has a mass number of 30
c) The particle has an atomic number of 20
d) The particle contains 20 neutrons

24. Which elements did Mendeleev leave spaces for in
his periodic table?
a) Be b) Sc

25. If copper metal is a mixture two isotopes, Cu-63,
mass = 62.994 g and Cu-65, mass = 63.925 g.
What is the mass of copper in 60.000 grams.
Calculate the "percentages" of the two isotopes of
copper. Show your work.

26. What is the formula of the ionic compound formed
between Mg and F?
a) MgF b) MgF₂
c) MgF₃ d) Mg₂F₃

27. What is the formula of the ionic compound formed
between Ca and F?
a) CaF₂ b) CaF₃
c) CaF₄ d) CaF₅

28. What is the name of the SiO₄⁴⁻ ion?
a) silicon b) silicon tetraoxide
c) silicon d) hydrogen silicon

29. _____

$$2.5(62.925)(x) + (64.9278)(1-x) = 64.546$$

$$x = 0.1911$$

$$\text{Cu-63} \rightarrow 19.11\%$$

$$\text{Cu-65} \rightarrow 80.89\%$$

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CH Cherryholmes

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Table of Contents Ap Bio Pogil Chapter 48 Answer Key

1. Understanding the eBook Ap Bio Pogil Chapter 48 Answer Key
 - The Rise of Digital Reading Ap Bio Pogil Chapter 48 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ap Bio Pogil Chapter 48 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ap Bio Pogil Chapter 48 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ap Bio Pogil Chapter 48 Answer Key
 - Personalized Recommendations
 - Ap Bio Pogil Chapter 48 Answer Key User Reviews and Ratings
 - Ap Bio Pogil Chapter 48 Answer Key and Bestseller Lists
5. Accessing Ap Bio Pogil Chapter 48 Answer Key Free and Paid eBooks

- Ap Bio Pogil Chapter 48 Answer Key Public Domain eBooks
- Ap Bio Pogil Chapter 48 Answer Key eBook Subscription Services
- Ap Bio Pogil Chapter 48 Answer Key Budget-Friendly Options

6. Navigating Ap Bio Pogil Chapter 48 Answer Key eBook Formats

- ePUB, PDF, MOBI, and More
- Ap Bio Pogil Chapter 48 Answer Key Compatibility with Devices
- Ap Bio Pogil Chapter 48 Answer Key Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Ap Bio Pogil Chapter 48 Answer Key
- Highlighting and Note-Taking Ap Bio Pogil Chapter 48 Answer Key
- Interactive Elements Ap Bio Pogil Chapter 48 Answer Key

8. Staying Engaged with Ap Bio Pogil Chapter 48 Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Ap Bio Pogil Chapter 48 Answer Key

9. Balancing eBooks and Physical Books Ap Bio Pogil Chapter 48 Answer Key

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Ap Bio Pogil Chapter 48 Answer Key

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Ap Bio Pogil Chapter 48 Answer Key

- Setting Reading Goals Ap Bio Pogil Chapter 48 Answer Key
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Ap Bio Pogil Chapter 48 Answer Key

- Fact-Checking eBook Content of Ap Bio Pogil Chapter 48 Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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