

# Chapter 2 Review

1. Correct properties are characteristics of metals.

Which property means that you can pound the substance into a ball?

- a) ductility  
b) malleability  
c) conductivity  
d) luster

2. Which of the following is a period?

- a) 2  
b) 10  
c) 5  
d) 18  
e) 16

3. Which of the following is a transition metal?

- a) Cl  
b) Fe  
c) P  
d) Ca  
e) C

4. Which of the following is an alkali metal?

- a) Mg  
b) Sr  
c) Rb  
d) Al  
e) Br

5. Which of the following is an lanthanide?

- a) Sr  
b) Ba  
c) Cs  
d) P  
e) W

6. Which element has the highest melting point?

- a) Pb  
b) Fe  
c) Cu  
d) Hg  
e) Mg

7. Which ions exist in the

- a) negative electrode  
b) positive electrode  
c) power source  
d) gas inside the tube

8. In a cathode ray tube, electrons are first created

- a) in a positively charged plate  
b) in a negatively charged plate  
c) in the vacuum  
d) in the glass tube

9. Listed below are the charges and masses of four particles. Which one will be deflected the least in a mass spectrometer?

- a) +0.2 amu  
b) +1.1 amu  
c) +1.1 amu  
d) +1.1 amu

10. In a million oil drop type experiments, the charge

on that all droplets (Coulombs) was found to be

- 1.60  $\times 10^{-19}$  C  
3.20  $\times 10^{-19}$  C  
4.80  $\times 10^{-19}$  C  
6.40  $\times 10^{-19}$  C  
8.00  $\times 10^{-19}$  C

What is the charge on the electron according to this experiment?

- a)  $1.1 \times 10^{-19}$  C  
b)  $3.2 \times 10^{-19}$  C  
c)  $4.8 \times 10^{-19}$  C  
d)  $6.4 \times 10^{-19}$  C  
e)  $8.0 \times 10^{-19}$  C

11. Pictured below is a column of the Rutherford experiment. Which scattered particles gives the best evidence for the nucleus?



- a) 1  
b) 2  
c) 3  
d) 4  
e) 5

12. Which of the following is an isotope of the element with 28 protons (or 28) and 32 neutrons (or 32)?

- a) Ni-58  
b) Ni-60  
c) Ni-62  
d) Ni-64

13. The imaginary element X has the following natural abundances and isotopic masses. What is the atomic mass of X?

- 25% 24.02 amu  
75% 26.10 amu

Show your work.

$$(24.02 \times 0.25) + (26.10 \times 0.75) =$$

$$25.3 \text{ amu}$$

For questions 14–17, use the following key:

(each answer may be used once, more than once,

or not at all)

- a) alpha  
b) beta  
c) gamma  
d) alpha and beta, but not gamma

14. A high energy form of light

15. Two protons &amp; two neutrons

16. A high speed electron

17. Used by Ernest Rutherford as a "probe"

a  
b  
c  
d  
e

For questions 18–22, use the following key:

(each answer may be used once, more than once,

or not at all)

- a) John Dalton  
b) Ernest Rutherford  
c) J.J. Thomson  
d) Democritus

18. His model of the atom has been called the "plum pudding" model

19. His model of the atom has been called the "billiard ball" model

20. His model of the atom has been called the "raisin bread" model

21. His philosophical idea included the term "atomos"

22. He added to the atomic theory the idea that atoms had positive and negative parts

a  
b  
c  
d  
e

 23. Consider the following isotopes:  $^{63}_{29}\text{Cu}$  and  $^{65}_{29}\text{Cu}$ . Which statement below is correct?

- a) The particle contains 63 protons  
b) The particle has a mass number of 94  
c) The particle has an atomic number of 120  
d) The particle contains 120 neutrons

24. Which elements did Mendeleev leave spaces for in his periodic table?

6a 6e 5c

 25. Copper metal is a mixture two isotopes,  $^{63}_{29}\text{Cu}$ , mass = 62.9296 u and  $^{65}_{29}\text{Cu}$ , mass = 64.9278 u. The atomic mass of copper is 63.546 u. Calculate the % abundance of the two isotopes of copper. Show your work.

26. What is the formula of the ionic compound formed between Mg and Br?

- a)  $\text{MgBr}$   
b)  $\text{Mg}_2\text{Br}$   
c)  $\text{Mg}_2\text{Br}_2$   
d)  $\text{MgBr}_2$

27. What is the formula of the ionic compound formed between Ca and P?

- a)  $\text{Ca}_3\text{P}_2$   
b)  $\text{CaP}$   
c)  $\text{Ca}_2\text{P}_3$   
d)  $\text{Ca}_3\text{P}$

 28. What is the name of the  $\text{Mg}^{2+}$  ion?

- a) sulfur  
b) sulfur dioxide  
c) hydrogen sulfur  
d) sulfur

$$25)(62.9296)(x) + (64.9278)(1-x) = 63.546$$

$$x = 0.1911$$

$$\text{Cu-63} \rightarrow 19.11\%$$

$$\text{Cu-65} \rightarrow 80.89\%$$

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**CH Cherryholmes**



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