

## Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

### Chapters 17 & 18

**Essential Skills/ State Standards:**

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)  
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)  
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CSS: 7d)  
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)  
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d\*, 8c)  
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)  
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

**Vocabulary**

- |                                                                                                                                                                           |                                                                                                                                                                               |                                                                                                                                                                                                                                                                                                |                                                                                                                                                                                  |
|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| <ul style="list-style-type: none"> <li>• Condensation</li> <li>• Freezing</li> <li>• Vaporization</li> <li>• Melting</li> <li>• Boiling</li> <li>• Sublimation</li> </ul> | <ul style="list-style-type: none"> <li>• Deposition</li> <li>• Heat</li> <li>• Temperature</li> <li>• Thermochemistry</li> <li>• Exothermic</li> <li>• Endothermic</li> </ul> | <ul style="list-style-type: none"> <li>• Calorimeter</li> <li>• Joule</li> <li>• Specific heat</li> <li>• <math>q = m \cdot c \cdot \Delta T</math></li> <li>• <math>q = m \cdot \Delta H_{\text{vaporization}}</math></li> <li>• <math>q = m \cdot \Delta H_{\text{fusion}}</math></li> </ul> | <ul style="list-style-type: none"> <li>• Heat of reaction</li> <li>• Le Chatelier's principle</li> <li>• Equilibrium</li> <li>• Activation Energy</li> <li>• Catalyst</li> </ul> |
|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|

**Book Assignments:**

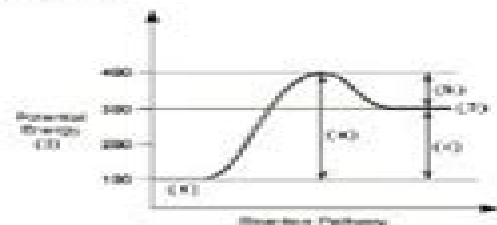
Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

**Study Guide (Use notes, worksheets, and labs in addition to this study guide)**

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:  

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$

a. Increase temp:	b. Increase HCl	c. Increase pressure	d. Remove $\text{MgCl}_2$	e. Add $\text{H}_2$
-------------------	-----------------	----------------------	---------------------------	---------------------
- Identify the following terms with their correct number/letter from the graphs:
  - \* Reactants- \_\_\_\_\_
  - \* Products- \_\_\_\_\_
  - \* Activation energy (number/letter & amount in Joules)- \_\_\_\_\_
  - \* Is this reaction exothermic or endothermic?
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



# Answers To Unit 8 Stoichiometry Study Guide

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*Fundamentals of Chemistry, Study Guide* James E. Brady,John R. Holum,1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs     **Study Guide** Joan I. Senyk,Darrell D. Ebbing,1984     General Chemistry Study Guide Sixth Edition Darrell D. Ebbing,1999     Study Guide for General Chemistry and College Chemistry, Eighth Editions by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw,1988     **McGraw-Hill Encyclopedia of Science and Technology** ,1977 Accompanied by Readers guide 33 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 1 Accompanied by Study guide 50 p 28 cm New York McGraw Hill c1977 Call number Q121 M3 1977 Suppl 2     **Laboratory Manual to Accompany Chemistry, [by] Stanley R. Radel, Marjorie H. Navidi** Arthur D. Baker,1990     **Chemistry, Study Guide** James E. Brady,John R. Holum,1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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