

2 AMU (Engineering) Solved Paper 2022

76. Camphor is often used in molecular mass determination because
 (a) it has a very high cryoscopic constant
 (b) it is volatile
 (c) it is insolvent for organic substances
 (d) it is readily available

77. Fe_3O_4 is ferromagnetic at room temperature but at 850 K it becomes
 (a) diamagnetic (b) paramagnetic
 (c) ferromagnetic (d) non-magnetic

78. To get n-type semiconductor, the impurity to be added to Si should have the following number of valence electrons
 (a) 1 (b) 2
 (c) 3 (d) 5

79. Melting process is used
 (a) for removal of Cu_2O from Cu
 (b) for removal of Al_2O_3 from Al
 (c) for removal of Fe_2O_3 from Fe
 (d) for removal of Cu_2O from Cu_2O_3

80. A gas present in a cylinder filled with a frictionless piston expands against a constant pressure of 1 atm from a volume of 2 litres to a volume of 6 litres. In doing so it absorbs 800 J heat from surroundings. Determine increase in internal energy of process
 (a) 400.05 J (b) 96.81 J
 (c) 394.05 J (d) 1205.04 J

81. Which statement is false?
 (a) van der Waals' constant 'a' is measure of magnitude of intermolecular attractive force with in the gas.
 (b) At low temperature, intermolecular attraction is more significant and van der Waals' constant 'a' is high at low temperature.
 (c) van der Waals' constant 'a' is independent of the pressure of the gas.
 (d) van der Waals' gas approximates ideal gas behaviour when pressure approaches zero.

82. The resistance of a 0.1 N solution of a salt is found to be 2.5×10^3 ohms. Calculate the equivalent conductance of the solution
 Cell constant = 1.15 cm^{-1}
 (a) $2.5 \times 10^3 \text{ ohm}^{-1} \text{ cm}^2 \text{ equiv}^{-1}$
 (b) $4.00 \text{ ohm}^{-1} \text{ cm}^2 \text{ equiv}^{-1}$
 (c) $5.00 \text{ ohm}^{-1} \text{ cm}^2 \text{ equiv}^{-1}$
 (d) $0.50 \text{ ohm}^{-1} \text{ cm}^2 \text{ equiv}^{-1}$

83. Boiling point
 (a) H_2O
 (b) H_2S
 (c) H_2Se
 (d) H_2Te

84. When sodium is dissolved in liquid NH_3 , a solution of deep blue colour is formed. The colour of the solution is due to
 (a) ammonium cation
 (b) sodium ion
 (c) sodium amide
 (d) ammonium sodium ion

85. The respective number of bridging 'CO' groups in $[\text{Co}_2(\text{CO})_8]$ and $[\text{Mn}_2(\text{CO})_{10}]$ is/are
 (a) 2, 0 (b) 1, 0 (c) 0, 1 (d) 2, 2

86. One mole of $\text{PbCl}_2 \cdot 4\text{NH}_3$ upon reaction with excess of AgNO_3 gives two moles of AgCl . The coordination number of complex formed will be
 (a) two (b) four (c) six (d) three

87. The correct electron gain enthalpy trends is
 (a) $\text{O} < \text{S} < \text{Se} < \text{Te} < \text{Po}$
 (b) $\text{O} > \text{S} > \text{Se} > \text{Te} > \text{Po}$
 (c) $\text{O} > \text{S} > \text{Se} > \text{Te} > \text{Po}$
 (d) $\text{O} < \text{S} > \text{Se} < \text{Te} > \text{Po}$

88. MnO_4^- is tetrahedral in geometry. There is π -bonding between Mn and oxygen by the overlap of
 (a) π - π orbitals (b) π - σ orbitals
 (c) π - δ orbitals (d) π - δ orbitals

89. The highest magnetic moment among the following coordinated compound is shown by
 (a) $\text{K}_4[\text{Mn}(\text{CN})_6]$ (b) $[\text{Fe}(\text{OH}_2\text{O})_6]^{2+}$
 (c) $\text{K}_4[\text{MnO}_4]$ (d) $\text{Ni}(\text{CO})_4$

90. The transition metal oxides or halides does not exist among the following is
 (a) V_2O_5 (b) MnF_3
 (c) CrF_3 (d) Mn_2O_7

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