

2. AMU (Engineering) Solved Paper 2022

10. Camphor is often used in molecular mass determination because
 (a) it has a very high cryoscopic constant
 (b) it is volatile
 (c) it is solvent for organic substances
 (d) it is readily available
11. Fe_2O_3 is ferromagnetic at room temperature but at 850 K it becomes
 (a) diamagnetic (b) paramagnetic
 (c) ferromagnetic (d) non-magnetic
12. To get n-type semiconductor, the impurity to be added to Si should have the following number of valence electrons
 (a) 1 (b) 2
 (c) 3 (d) 5
13. Pilling process is used
 (a) for removal of Cu_2O from Cu
 (b) for removal of Al_2O_3 from Al
 (c) for removal of Fe_2O_3 from Fe
 (d) for removal of Cu_2O from Cu_2O_2
14. A gas present in a cylinder filled with a frictionless piston expands against a constant pressure of 1 atm from a volume of 2 litres to a volume of 6 litres. In doing so it absorbs 800 J heat from surroundings. Determine increase in internal energy of process
 (a) 405.04 J (b) 95.81 J
 (c) 394.95 J (d) 1205.04 J
15. Which statement is false?
 (a) van der Waals' constant 'a' is measure of magnitude of intermolecular attractive force with in the gas.
 (b) At low temperature, intermolecular attraction is more significant and van der Waals' constant 'a' is high at low temperature.
 (c) van der Waals' constant 'a' is independent of the pressure of the gas.
 (d) van der Waals' gas approximates ideal gas behaviour when pressure approaches zero.
16. The resistance of a 0.1 N solution of a salt is found to be 2.5×10^3 ohms. Calculate the equivalent conductance of the solution
 Cell constant = 1.15 cm^{-1}
 (a) $2.30 \text{ ohm}^{-1} \text{ cm}^2 \text{ eqvt}^{-1}$
 (b) $4.60 \text{ ohm}^{-1} \text{ cm}^2 \text{ eqvt}^{-1}$
 (c) $5.60 \text{ ohm}^{-1} \text{ cm}^2 \text{ eqvt}^{-1}$
 (d) $0.90 \text{ ohm}^{-1} \text{ cm}^2 \text{ eqvt}^{-1}$
17. Borazine is
 (a) $\text{B}_2\text{H}_2\text{H}_6$ (b) B_2H_4
 (c) $\text{B}_2\text{H}_2\text{H}_6$ (d) $\text{B}_2\text{H}_2\text{H}_6$
18. When sodium is dissolved in liquid NH_3 , a solution of deep blue colour is formed. The colour of the solution is due to
 (a) ammoniated electron
 (b) sodium ion
 (c) sodium amide
 (d) ammoniated sodium ion
19. The respective number of bridging 'CO' groups in $[\text{Co}_2(\text{CO})_8]$ and $[\text{Mn}_2(\text{CO})_{10}]$ is/are
 (a) 2, 0 (b) 1, 0 (c) 0, 1 (d) 2, 2
20. One mole of $\text{PbCl}_2 \cdot 4\text{NH}_3$ upon reaction with excess of AgNO_3 gives two moles of AgCl . The coordination number of complex formed will be
 (a) two (b) four (c) six (d) three
21. The correct electron gain enthalpy trends is
 (a) $\text{O} < \text{S} < \text{Se} < \text{Te} < \text{Po}$
 (b) $\text{O} < \text{S} > \text{Se} > \text{Te} > \text{Po}$
 (c) $\text{O} > \text{S} > \text{Se} > \text{Te} > \text{Po}$
 (d) $\text{O} < \text{S} > \text{Se} < \text{Te} > \text{Po}$
22. MnO_4^- is tetrahedral in geometry. There is π -bonding between Mn and oxygen by the overlap of
 (a) p - p orbitals (b) s - p orbitals
 (c) p - d orbitals (d) s - d orbitals
23. The highest magnetic moment among the following coordinated compound is shown by
 (a) $\text{K}_2[\text{Mn}(\text{CN})_6]$ (b) $[\text{Fe}(\text{H}_2\text{O})_6]^{2+}$
 (c) $\text{K}_2[\text{MnO}_4]$ (d) $\text{Ni}(\text{CO})_4$
24. The transition metal oxides or halides does not exist among the following is
 (a) V_2O_5 (b) MnF_5
 (c) CrF_6 (d) Mn_2O_7

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Tao Wei



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