

Experiment 7

Advance Study Assignment: Analysis of an Unknown Chloride

1. A student performed this experiment and obtained the following concentration values: 0.02813 M, 0.02802 M, and 0.02788 M.

a. What is the mean concentration?

— 10 —

b. What is the standard deviation of these results?

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2. A 10.00 mL-diluted chloride sample required 13.89 mL of 0.02014 M AgNO_3 to reach the Fajans endpoint.

a. How many moles of Cl^- ions were present in the sample? (Use Eqs. 2 and 3.)

$$\begin{aligned}
 \text{moles of } \text{Ag}^+ &= \text{Moles of } \text{AgNO}_3 \\
 &= 0.02014 \text{ M} \times 0.01389 \text{ L} \\
 &= 0.00028 \text{ moles}
 \end{aligned}
 \quad \text{0.00028 moles of }$$

b. What was the concentration of chloride in the diluted solution?

$$M_2 = \frac{0.02014 \text{ M} \times 13.89 \text{ mL}}{10.00 \text{ mL}} = \underline{0.028 \text{ M}}$$

3. How would the following errors affect the concentration of Cl^- obtained in Question 2b? Give your reasoning in each case.

a. The student read the molarity of AgNO_3 as 0.02104 M instead of 0.02014 M.

b. The student was past the endpoint of the situation when he took the final measurement.

Analysis Of An Unknown Chloride Answers

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Environmental Sampling and Analysis Maria Csuros,1997-03-24 This manual covers the latest laboratory techniques state of the art instrumentation laboratory safety and quality assurance and quality control requirements In addition to complete coverage of laboratory techniques it also provides an introduction to the inorganic nonmetallic constituents in environmental samples their chemistry and their control by regulations and standards Environmental Sampling and Analysis Laboratory Manual is perfect for college and graduate students learning laboratory practices as well as consultants and regulators who make evaluations and quality control decisions Anyone performing laboratory procedures in an environmental lab will appreciate this unique and valuable text

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Analytical Chemistry Gary D. Christian,Purnendu K. Dasgupta,Kevin A. Schug,2013-10-07 With the 7th Edition of Analytical Chemistry renowned chemists Purnendu Sandy Dasgupta and Kevin Schug both of the University of Texas Arlington join the author team The new edition focuses on more in depth coverage of the principles and techniques of quantitative analysis and instrumental analysis aka Analytical Chemistry The goal of the text is to provide a foundation of the analytical process tools and computational methods and resources and to illustrate with problems that bring realism to the practice and importance of analytical chemistry It is designed for undergraduate college students majoring in chemistry and in fields related to chemistry

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Quantitative Analysis William Rieman,Jacob David Neuss,Barnet Naiman,1951 Common apparatus and operations The balance Volumetric apparatus Fundamentals of volumetric analysis Volumetric determination of chloride ion by Mohr's method Potentiometric measurements Ionization of salts acids and bases Acidimetry and alkalimetry

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News and Journal of Physical Science ,1900 The American Chemist Charles Frederick Chandler,William Henry Chandler,1871 American contributions to Chemistry By Benjamin Silliman v 5 p 70 114 195 209 *Qualitative Analysis and Ionic Equilibrium* George H. Schenk,Darrell D. Ebbing,1985 A supplement for courses with a qualitative analysis component this lab manual contains explanations of the chemistry of metal ions and anions It includes pre lab exercises experiments and lab reports *The Chemical Review* ,1873 English Mechanic and Mirror of Science ,1876

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