

ACS Practice Test 1

Acids & Bases:

1. All are examples of Lewis acid–base reactions *except*

- (A) $\text{Cu}^{2+}(\text{aq}) + 4\text{NH}_3(\text{aq}) \rightleftharpoons [\text{Cu}(\text{NH}_3)_4]^{2+}(\text{aq})$
 (B) $\text{HCl}_{(\text{g})} + \text{NH}_{3(\text{g})} \rightarrow \text{NH}_4\text{Cl}_{(\text{s})}$
 (C) $\text{H}^{+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightleftharpoons \text{H}_2\text{O}_{(\text{l})}$
 (D) $2\text{Na}_{(\text{s})} + \text{Cl}_{2(\text{g})} \rightarrow 2\text{NaCl}_{(\text{s})}$

2. According to the Lewis definition, an acid is a species

- (A) having a hydrogen ion. (B) donating a pair of electrons.
 (C) accepting a pair of electrons. (D) accepting a hydrogen ion.

3. All are potential Lewis bases *except*

- (A) NH_3 (B) H_2O (C) CH_4 (D) CN^{-}

4. Which hydrated metal ion is most acidic under conditions of equal molar concentration in water?

- (A) Al^{3+} (B) Ba^{2+} (C) K^{+} (D) Zn^{2+} (E) Ag^{+}

5. Which element in Group 5A forms the most basic oxide?

- (A) N (B) P (C) As (D) Sb (E) Bi

Atomic Structure:

6. In what respect does an atom of magnesium differ from a magnesium ion (Mg^{2+})?

- (A) The ion has a more stable electronic arrangement than the atom.
 (B) The positive charge on the nucleus of the ion is two units greater than the nuclear charge on the atom.
 (C) The ion has two more protons than the atom.
 (D) The ion will react more readily with free Cl_2 .
 (E) The ion has two more planetary electrons than the atom.

7. A certain atom has the symbol $^{148}_{73}\text{Pc}$. What does this symbol tell about an atom of the element?

- (A) It has 148 electrons. (B) It has a valence of 4.
 (C) It has 221 protons in the nucleus. (D) It has 75 neutrons in the nucleus.
 (E) Its atomic mass is 73 u.

8. The atomic mass of an element is 32.07 u and its atomic number is 16. The number of protons contained in the nucleus of an atom of this element is

- (A) 16 (B) 31 (C) 32 (D) 48

9. The atomic mass of an element is 32.07 u and its atomic number is 16. The element forms a simple ion. The ion will most likely have a charge of

- (A) 1− (B) 2− (C) 3+ (D) 1+

Adlc Chemistry 30 Answer Key

Daniel F McAuley



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