

## 19.2

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $K_a$  or  $K_b$
2. Calculate  $K_a$  and  $K_b$  for weak acids and bases, and vice versa, given  $pK_a$  or  $pK_b$
3. Calculate the percent ionization of weak acids

## Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant ( $K_a$ )
- acid ionization

- base dissociation
- base ionization

## Key Equations

$$K_a = \frac{[\text{H}^+][\text{A}^-]}{[\text{HA}]} \quad \text{Equation 19.1}$$

$$K_b = \frac{[\text{B}][\text{OH}^-]}{[\text{BOH}]} \quad \text{Equation 19.2}$$

## Part A Completion

Write the equilibrium constant for the acid ionization of hydrofluoric acid and water. Identify the products and reactants in the equilibrium.

Hydrofluoric acid is a weak acid. Write the equilibrium constant for the acid ionization of hydrofluoric acid.

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# 19 2 Hydrogen Ions And Acidity Section Review

**Christian Drosten**



## **19 2 Hydrogen Ions And Acidity Section Review:**

## **19 2 Hydrogen Ions And Acidity Section Review** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**19 2 Hydrogen Ions And Acidity Section Review**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve to the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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