

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization constant | 5. K_a | 6. K_b |
| | 7. weak solution | |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes. K_w is the ionization constant of water. Write the equation and K_w at 25°C.
2. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
3. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
4. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
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7. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
8. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
9. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.
10. The autoionization constant of water is 1.0×10^{-14} at 25°C. Write the equation and K_w at 25°C.

1. $K_w = 1.0 \times 10^{-14}$ 2. $K_w = 1.0 \times 10^{-14}$ 3. $K_w = 1.0 \times 10^{-14}$ 4. $K_w = 1.0 \times 10^{-14}$ 5. $K_w = 1.0 \times 10^{-14}$ 6. $K_w = 1.0 \times 10^{-14}$ 7. $K_w = 1.0 \times 10^{-14}$ 8. $K_w = 1.0 \times 10^{-14}$ 9. $K_w = 1.0 \times 10^{-14}$ 10. $K_w = 1.0 \times 10^{-14}$ 11. $K_w = 1.0 \times 10^{-14}$ 12. $K_w = 1.0 \times 10^{-14}$ 13. $K_w = 1.0 \times 10^{-14}$ 14. $K_w = 1.0 \times 10^{-14}$ 15. $K_w = 1.0 \times 10^{-14}$ 16. $K_w = 1.0 \times 10^{-14}$ 17. $K_w = 1.0 \times 10^{-14}$ 18. $K_w = 1.0 \times 10^{-14}$ 19. $K_w = 1.0 \times 10^{-14}$ 20. $K_w = 1.0 \times 10^{-14}$

19 2 Hydrogen Ions And Acidity Worksheet Answers

Hussin A.Rothana

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Reviewing **19 2 Hydrogen Ions And Acidity Worksheet Answers**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**19 2 Hydrogen Ions And Acidity Worksheet Answers**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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