

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for a given pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pH = -\log[H_3O^+]$ |

Part A: Completion

Use the appropriate number to indicate the correctness of the response provided that corresponds to the correct response. Circle the number in the space provided.

1. When a substance is _____, it is an acid. Hydrogen ion concentration is _____ in the solutions of these acids prepared in water. _____.
2. _____ is the concentration of _____ in a solution. _____.
3. _____ is the concentration of _____ in a solution. _____.
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Advisory Council on College Chemistry

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