

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and vice versa, given K_a or K_b .
2. Calculate K_a and K_b for weak acids and bases, and vice versa, given pK_a or pK_b .
3. Calculate the percent ionization of weak acids.

Vocabulary

- acid dissociation
- acid ionization

- acid ionization constant (K_a)
- acid ionization

- base ionization
- base ionization constant (K_b)

Key Equations

- 1. $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1)

- 2. $K_b = \frac{[HB^+][OH^-]}{[B]}$ (Equation 19.2)

Part A Completion

Write the equilibrium constant for the acid ionization of hydrofluoric acid and water. Identify the products and reactants, and write the equilibrium constant expression for the reaction.

Hydrofluoric acid is a weak acid. Write the equilibrium constant expression for the reaction.

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1. $K_a = \frac{[H^+][F^-]}{[HF]}$
2. $K_b = \frac{[HF][OH^-]}{[F^-]}$
3. $K_a = \frac{[H^+][F^-]}{[HF]}$
4. $K_b = \frac{[HF][OH^-]}{[F^-]}$
5. $K_a = \frac{[H^+][F^-]}{[HF]}$
6. $K_b = \frac{[HF][OH^-]}{[F^-]}$
7. $K_a = \frac{[H^+][F^-]}{[HF]}$
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9. $K_a = \frac{[H^+][F^-]}{[HF]}$
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17. $K_a = \frac{[H^+][F^-]}{[HF]}$
18. $K_b = \frac{[HF][OH^-]}{[F^-]}$
19. $K_a = \frac{[H^+][F^-]}{[HF]}$
20. $K_b = \frac{[HF][OH^-]}{[F^-]}$

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Sabine Zange



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