

THERMODYNAMICS

EXERCISE - I

SINGLE CORRECT (OBJECTIVE QUESTIONS)

Sol.1 [C]

Sol.2 [D]

For 3 moles of Cl_2 & HCl
 $\Delta U = -3 \times 185 = -555 \text{ kJ}$

Sol.3 [D]

2 mole HCl is consumed completely
 for 1 mole consumption of HCl
 $\Delta U = \Delta H - \Delta n_g \cdot RT$
 $= -72.3 + \frac{8.314 \times 300}{1000} \text{ (as, } \Delta n_g = -1\text{)}$
 $= -69.8 \text{ kJ}$

For 2 mole HCl
 $\Delta U = 2 \times -69.8 = -139.6 \text{ kJ}$

Sol.4 [B]

$$\Delta U = Q + W$$

$$= \frac{100 \times 10^3 - 1.5 \times 10^5}{10^3} = -50 \text{ kJ}$$

Sol.5 [C]

$$w = nC_v \cdot \Delta T$$

$$\Rightarrow -3 \times 1000 = 1 \times 20 \times (T - 300)$$

$$T = 150 \text{ K}$$

Sol.6 $\Delta U = Q + w$

$$Q = -124 \text{ J}$$

$$W = -P_{ext}(v_2 - v_1)$$

$$= -\frac{1520}{760} (177 - 377) = 400 \text{ Atm ml} = 40 \text{ J}$$

$$\Delta U = -124 + 40 = -84 \text{ J}$$

Sol.7 [C]

Heat released by 500 gm of liquid

$$= \frac{500}{18} \times 75.6 \times 20$$

Heat absorbed by 'n' no of ice each 9 gm

$$= \frac{9}{12} \cdot n \times 6000 \text{ J}$$

Heat absorbed = Heat released
 Which gives $n = 14$

Sol.8 [A]

Area of $\Delta ABC = -$ Area of ΔACD

Sol.9 [B]

$$C_p = \frac{1}{n} \left(\frac{dq}{dT} \right)_p$$

$$dT = 0$$

Sol.10 [C]

Sol.11 [D]

$$nC_v \cdot \Delta T = 100 \times 4.2 = 209$$

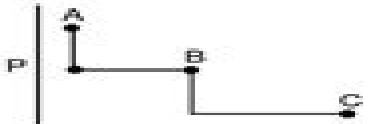
$$= 420 - 209 = 211 \text{ J}$$

$$n \times \frac{5R}{2} \times \Delta T = 211$$

$$n\Delta T = \frac{211 \times 2}{5R}$$

$$C_m = \frac{q}{n\Delta T} = \frac{420}{211 \times 2} \times 5R = 5R$$

Sol.12 [C]



$$\text{Total work done} = -2(16 - 4) - 1(32 - 16)$$

$$= -24 - 16 = -40 \text{ bar litre}$$

$$= -4000 \text{ J}$$

$$Q = -W = 4000 \text{ J}$$

Sol.13 [A]

Sol.14 [C]

$$y = \frac{7}{5}$$

$$y - 1 = \frac{2}{5}$$

$$T_2 = T_1 \left(\frac{y_1}{y_2} \right)^{\frac{2}{5}}$$

$$= 600 \times \left(\frac{1}{32} \right)^{2/5} = 150 \text{ K}$$

$$\therefore \Delta H = nC_p(T_2 - T_1) = -\frac{7}{2} R \times 450 = -1575 \text{ R}$$

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