

Name _____
Specific Heat Worksheet

Key

$C_p = q/m\Delta T$, where q = heat energy, m = mass, and T = temperature

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the heat capacity of iron.

$$1086.75 = 15.75g \cdot 150 \cdot C_p$$

$$1086.75 = 2362.5 C_p$$

$$(C_p = .46 \text{ J/g°C})$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = C_p \cdot m \cdot \Delta T$$

$$.90(10)(33) = [297 \text{ J}]$$

$$\frac{55}{22} = \frac{33}{33}$$

3. To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 joules of heat and its heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20.0°C.

$$5275 = .50(50)\Delta T$$

$$\Delta T = 211$$

$$T_f = +\frac{211}{20} [231^\circ\text{C}]$$

4. Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs 6.75×10^4 joules of heat, and its temperature changes from 32°C to 57°C.

$$67500 = C_p \cdot 1500 \cdot 25^\circ\text{C}$$

$$67500 = 37500 C_p$$

$$[1.8 \text{ J/g°C}]$$

5. 100.0 mL of 4°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat needed to cause this rise in temperature.

$$Q = 100g \cdot 4.18 \cdot 33^\circ\text{C}$$

$$= 13794$$

$$[13794]$$

6. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the heat capacity of mercury.

$$455 = C_p \cdot 25g \cdot 130$$

$$455 = 3250 C_p$$

$$-.14 = C_p$$

7. What is the specific heat capacity of silver metal if 55.00 g of the metal absorbs 198 joules of heat and the temperature rises 15.0°C?

$$198 = C_p \cdot 55g \cdot 15^\circ\text{C}$$

$$198 = 825 C_p$$

$$C_p = .24 \text{ J/g°C}$$

8. What is the change in temperature of 150.0 g of chloroform if it absorbs 1000.0 joules of heat, and the specific heat of chloroform is 0.96 J/g°C?

$$1000 = .96(150g)\Delta T$$

$$1000 = 144 \Delta T$$

$$6.94 = \Delta T$$

$$[6.9^\circ\text{C}]$$

Answers:

1. $C_p = .46 \text{ J/g°C}$
2. $Q = 297 \text{ J}$
3. $T_{final} = 231^\circ\text{C}$
4. $C_p = 1.8 \text{ J/g°C}$
5. $Q = 13800 \text{ J}$ **13794**
6. $C_p = .14 \text{ J/g°C}$
7. $C_p = .24 \text{ J/g°C}$
8. $\Delta T = 6.9^\circ\text{C}$

Answer Key For Specific Heat

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Table of Contents Answer Key For Specific Heat

1. Understanding the eBook Answer Key For Specific Heat
 - The Rise of Digital Reading Answer Key For Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Specific Heat
 - Personalized Recommendations
 - Answer Key For Specific Heat User Reviews and Ratings
 - Answer Key For Specific Heat and Bestseller Lists
5. Accessing Answer Key For Specific Heat Free and Paid eBooks
 - Answer Key For Specific Heat Public Domain eBooks
 - Answer Key For Specific Heat eBook Subscription Services
 - Answer Key For Specific Heat Budget-Friendly Options
6. Navigating Answer Key For Specific Heat eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Specific Heat Compatibility with Devices
 - Answer Key For Specific Heat Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Specific Heat
 - Highlighting and Note-Taking Answer Key For Specific Heat
 - Interactive Elements Answer Key For Specific Heat
8. Staying Engaged with Answer Key For Specific Heat

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Key For Specific Heat

9. Balancing eBooks and Physical Books Answer Key For Specific Heat

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answer Key For Specific Heat

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answer Key For Specific Heat

- Setting Reading Goals Answer Key For Specific Heat
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answer Key For Specific Heat

- Fact-Checking eBook Content of Answer Key For Specific Heat
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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