

Name: _____ Date: _____

Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

1. The Latin word *color* means “heat,” and meter comes from the Greek word meaning “to measure.” What do you think a calorimeter does? **Measures heat energy in an object**
2. Where have you heard the word calorie before? What do you think a calorie is? **A measure of heat energy**

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to find a substance’s specific heat capacity. You will use the Calorimetry Lab Gizmo to determine the specific heat capacities of various substances.

1. On the **SIMULATION** pane, select Copper. Use the slider to set its Mass to 200 g. Set the Water mass to 200 g. Check that the Water temp is set to 30.0 °C and the copper’s Temp is 90 °C. Select the **GRAPH** tab, and click **Play**.

A. What was the Final temperature of the copper and the water? **34.96 °C**

B. How much did the temperature of the copper change? **-55.04 °C**

C. How much did the temperature of the water change? **4.96 °C**

2. Specific heat capacity can be described as a substance’s resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain. **Water, because the temperature of the water did not change that much. It would have taken more energy to raise the temperature of the water any higher than it already is.**

Activity A:

Heat transfer

Get the Gizmo ready:

Click **Reset**.

Question: What factors determine how heat energy transfers between objects?

1. Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g of 30.0 °C water.

A. How do you think increasing the water’s mass would affect the final temperature? **It would lower the temperature of the water a little bit**

B. How do you think decreasing the copper’s mass would affect the final temperature? **It would increase the temperature of the water a little bit**

C. How do you think increasing or decreasing the copper’s initial temperature would affect the final temperature? **There would be a slight raise in temperature**

Answer Key Explore Learning Calorimetry Lab

Jerry D. Wilson



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College Physics Jerry D. Wilson,2000 **Electrical & Electronics Abstracts** ,1989 **Government Reports**
Announcements & Index ,1994-07 The Software Encyclopedia 2001 ,2001 **The Software Encyclopedia 2000**
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Table of Contents Answer Key Explore Learning Calorimetry Lab

1. Understanding the eBook Answer Key Explore Learning Calorimetry Lab
 - The Rise of Digital Reading Answer Key Explore Learning Calorimetry Lab
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Explore Learning Calorimetry Lab
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Explore Learning Calorimetry Lab
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Explore Learning Calorimetry Lab
 - Personalized Recommendations
 - Answer Key Explore Learning Calorimetry Lab User Reviews and Ratings
 - Answer Key Explore Learning Calorimetry Lab and Bestseller Lists
5. Accessing Answer Key Explore Learning Calorimetry Lab Free and Paid eBooks
 - Answer Key Explore Learning Calorimetry Lab Public Domain eBooks
 - Answer Key Explore Learning Calorimetry Lab eBook Subscription Services
 - Answer Key Explore Learning Calorimetry Lab Budget-Friendly Options
6. Navigating Answer Key Explore Learning Calorimetry Lab eBook Formats

- ePub, PDF, MOBI, and More
 - Answer Key Explore Learning Calorimetry Lab Compatibility with Devices
 - Answer Key Explore Learning Calorimetry Lab Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answer Key Explore Learning Calorimetry Lab
 - Highlighting and Note-Taking Answer Key Explore Learning Calorimetry Lab
 - Interactive Elements Answer Key Explore Learning Calorimetry Lab
8. Staying Engaged with Answer Key Explore Learning Calorimetry Lab
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Explore Learning Calorimetry Lab
9. Balancing eBooks and Physical Books Answer Key Explore Learning Calorimetry Lab
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Explore Learning Calorimetry Lab
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key Explore Learning Calorimetry Lab
- Setting Reading Goals Answer Key Explore Learning Calorimetry Lab
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key Explore Learning Calorimetry Lab
- Fact-Checking eBook Content of Answer Key Explore Learning Calorimetry Lab
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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