

**15.1****ELECTRON CONFIGURATION IN IONIC BONDING**  
**SECTION REVIEW****Objectives**

- Use the periodic table to determine the number of valence electrons for an atom and draw its electron-dot structure.
- Describe the bonding and ionic radius of ionic compounds and ionic bonds in terms of electron transfer.

**Key Terms**

- valence electrons
- electron-dot structure
- ionic bond
- ion

**Part A: Completion**

After this completion exercise, you will know and understand all the concepts about ionic bonding and ionic structures. Each blank has the completed word or term, where possible, written here.

Elements within the same group of the periodic table have \_\_\_\_\_ (1) \_\_\_\_\_.

Similarly, elements in the same period have \_\_\_\_\_ (2) \_\_\_\_\_.

\_\_\_\_\_ (3) \_\_\_\_\_ is a set of 8 valence electrons called a noble gas.

Electrons in the outermost shell of an atom are called the \_\_\_\_\_ (4) \_\_\_\_\_ of the atom.

Electrons within the same group of the periodic table have \_\_\_\_\_ (5) \_\_\_\_\_.

Atoms are held together by \_\_\_\_\_ (6) \_\_\_\_\_ forces.

Attraction of valence electrons to a positively charged ion is called \_\_\_\_\_ (7) \_\_\_\_\_.

Attraction of valence electrons to a negatively charged ion is called \_\_\_\_\_ (8) \_\_\_\_\_.

The valence electrons of a noble gas are called the \_\_\_\_\_ (9) \_\_\_\_\_ of the noble gas.

Attraction of valence electrons to a positively charged ion is called \_\_\_\_\_ (10) \_\_\_\_\_.

Attraction of valence electrons to a negatively charged ion is called \_\_\_\_\_ (11) \_\_\_\_\_.

Attraction of valence electrons to a positively charged ion is called \_\_\_\_\_ (12) \_\_\_\_\_.

**Part B: True-False**

Classify each of these statements as true or false. A T is a true statement, an F is a false statement, and a NC is an answer that is not clear.

\_\_\_\_\_ (13) \_\_\_\_\_ The valence shell gains or loses electrons to become an ion.

\_\_\_\_\_ (14) \_\_\_\_\_ The valence shell gains or loses electrons to become a noble gas.

# Answer Key Electron Configuration In Ionic Bonding

**Disha Experts**

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