

Periodic Trends

Can the properties of an element be predicted using a periodic table?

Why?

The periodic table is often considered to be the “best friend” of chemists and chemistry students alike. It includes information about atomic masses and element symbols, but it can also be used to make predictions about atomic size, electronegativity, ionization energies, bonding, solubility, and reactivity. In this activity you will look at a few periodic trends that can help you make those predictions. Like most trends, they are not perfect, but useful just the same.

1. Consider the data in Model 1 on the following page.
 - a. Each element has three numbers listed under it. Which value represents the atomic radius? *The first number shows below the electron shell diagram.*
 - b. What are the units for the atomic radius? *pm—picometers*
 - c. Write a complete sentence to convey your understanding of atomic radius. Hint: You *may* not use the word “radius” in your definition.
Answers will vary. The distance from the center of the atom (nucleus) to the outer electrons.
2. In general, what is the trend in atomic radius as you go down a group in Model 1? Support your answer, using examples from three groups.
The atomic radius gets larger as you go down a group (carbon) in Model 1. Thus, Li is smaller than Be, B is smaller than Al, F is smaller than Cl, etc.
3. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 2. Hint: You should discuss either a change in distance between the nucleus and outer shell of electrons or a change in the number of protons in the nucleus.
Atomic radii increase top to bottom because the number of energy levels increases, increasing the “pull” on electrons from the protons in the nucleus and making the size of the atom larger. The highest occupied energy level has the greatest radius because as the number of electrons increases, more energy levels need to be added further away from the nucleus, making the size of the atom increase.
4. In general, what is the trend in atomic radius as you go across a period (left to right) in Model 1? Support your answer, using examples from two periods.
The atomic radius decreases from left to right across a period. Li is larger than F, Ne is larger than Cl, etc.
5. Using your knowledge of Coulombic attraction and the structure of the atom, explain the trend in atomic radius that you identified in Question 4.
Atomic radii decrease left to right because the number of protons increases, which increases the attractive force on the electrons, and pulls the electrons in the same energy level closer to the nucleus.

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