

Balancing Chemical Reactions

Balance all of the chemical reactions below. All compounds are written correctly.

After balancing the reaction, determine the type(s) of chemical reaction(s).

TYPE(S) of RXN

1. 2 $\text{NH}_3 \rightarrow$ _____ $\text{N}_2 +$ 3 H_2 Decomp
2. 4 $\text{Al} +$ 3 $\text{O}_2 \rightarrow$ 2 Al_2O_3 Syn/Comb
3. _____ $\text{N}_2\text{H}_4 +$ _____ $\text{O}_2 \rightarrow$ _____ $\text{N}_2 +$ 2 H_2O S.R. (?)
4. 2 $\text{Al} +$ 3 $\text{H}_2\text{SO}_4 \rightarrow$ 3 $\text{H}_2 +$ _____ $\text{Al}_2(\text{SO}_4)_3$ S.R.
5. _____ $\text{P}_2\text{O}_5 +$ 3 $\text{H}_2\text{O} \rightarrow$ 2 H_3PO_4 Syn
6. 2 $\text{C}_2\text{H}_2 +$ 5 $\text{O}_2 \rightarrow$ 4 $\text{CO}_2 +$ 2 H_2O Comb.
7. _____ $\text{Fe}_2\text{O}_3 +$ 3 $\text{H}_2 \rightarrow$ 2 $\text{Fe} +$ 3 H_2O S.R.
8. 4 $\text{FeS} +$ 7 $\text{O}_2 \rightarrow$ 2 $\text{Fe}_2\text{O}_3 +$ 4 SO_2 Comb.
9. 1 $\text{Cu}_2\text{S} +$ 3 $\text{O}_2 \rightarrow$ 2 $\text{Cu}_2\text{O} +$ 2 SO_2 Comb.
10. 4 $\text{NH}_3 +$ 5 $\text{O}_2 \rightarrow$ 4 $\text{NO} +$ 6 H_2O Comb.
11. _____ $\text{C}_3\text{H}_8 +$ 5 $\text{O}_2 \rightarrow$ 3 $\text{CO}_2 +$ 4 H_2O Comb
12. 2 $\text{AlPO}_4 +$ 3 $\text{Na}_2\text{S} \rightarrow$ _____ $\text{Al}_2\text{S}_3 +$ 2 Na_3PO_4 D.R.
13. _____ $\text{C}_{25}\text{H}_{52} +$ 38 $\text{O}_2 \rightarrow$ 26 $\text{H}_2\text{O} +$ 25 CO_2 Comb
14. 2 $\text{FeCl}_3 +$ 3 $\text{Na}_2\text{CO}_3 \rightarrow$ _____ $\text{Fe}_2(\text{CO}_3)_3 +$ 6 NaCl D.R.
15. _____ $\text{P}_4\text{O}_{10} +$ 6 $\text{H}_2\text{O} \rightarrow$ 4 H_3PO_4 Syn.
16. 4 $\text{NH}_3 +$ 7 $\text{O}_2 \rightarrow$ 4 $\text{NO}_2 +$ 6 H_2O Comb.
17. 2 $\text{C}_7\text{H}_8\text{O}_2 +$ 15 $\text{O}_2 \rightarrow$ 14 $\text{CO}_2 +$ 6 H_2O Comb.

Answer Key To Classifying Chemical Reactions

Michael J. Padilla



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