

1. As a solute is added to a solvent, what happens to the freezing point and the boiling point of the solution?

- (A) The freezing point decreases and the boiling point increases.  
 (B) The freezing point increases and the boiling point increases.  
 (C) The freezing point increases and the boiling point decreases.  
 (D) The freezing point decreases and the boiling point decreases.

- Note*  
 1.0 M  
 (A) lower (B) higher (C) the same

3. Which aqueous solution of KI freezes at the lowest temperature?

- (A) 1 mol of KI in 500. g of water  
 (B) 2 mol of KI in 500. g of water *(most concentrated)*  
 (C) 1 mol of KI in 1000. g of water  
 (D) 2 mol of KI in 1000. g of water

4. A 1 kilogram sample of water will have the highest freezing point when it contains

- (A)  $1 \times 10^{21}$  dissolved particles  
 (B)  $1 \times 10^{17}$  dissolved particles  
 (C)  $1 \times 10^{13}$  dissolved particles  
 (D)  $1 \times 10^{23}$  dissolved particles *(most dilute f.p.)*

5. A solution consists of 0.50 mole of  $\text{CaCl}_2$  dissolved in 100. grams of  $\text{H}_2\text{O}$  at  $25^\circ\text{C}$ . Compared to the boiling point and freezing point of 100. grams of  $\text{H}_2\text{O}$  at standard pressure, the solution at standard pressure has

- (A) a higher boiling point and a lower freezing point  
 (B) a lower boiling point and a lower freezing point  
 (C) a higher boiling point and a higher freezing point  
 (D) a lower boiling point and a higher freezing point

6. Which solution will freeze at the lowest temperature?

- (A) 1 mole of sugar in 500 g of water  
 (B) 1 mole of sugar in 1,000 g of water  
 (C) 2 moles of sugar in 1,000 g of water  
 (D) 2 moles of sugar in 500 g of water

7. Which solution has the highest boiling point at standard pressure?

- (A) 0.10 M  $\text{K}_2\text{PO}_4$  (aq) *(4 ions 3  $\text{K}^+$  & 1  $\text{PO}_4^{2-}$ )*  
 (B) 0.10 M  $\text{KCl}$  (aq)  
 (C) 0.10 M  $\text{KNO}_3$  (aq)  
 (D) 0.10 M  $\text{K}_2\text{SO}_4$  (aq)

8. Which concentration of a solution of  $\text{CH}_3\text{OH}$  in water has the lowest freezing point?

- (A) 0.001 M (B) 0.1 M  
 (C) 0.01 M (D) 0.0001 M

9. How do the boiling point and freezing point of a solution of water and calcium chloride at standard pressure compare to the boiling point and freezing point of water at standard pressure?

- (A) Both the freezing point and boiling point of the solution are higher.  
 (B) The freezing point of the solution is higher and the boiling point of the solution is lower.  
 (C) The freezing point of the solution is lower and the boiling point of the solution is higher.  
 (D) Both the freezing point and boiling point of the solution are lower.

10. Compared to a 2.0 M aqueous solution of  $\text{NaCl}$  at 1 atmosphere, a 3.0 M aqueous solution of  $\text{NaCl}$  at 1 atmosphere has a

- (A) lower boiling point and a higher freezing point  
 (B) higher boiling point and a higher freezing point  
 (C) lower boiling point and a lower freezing point  
 (D) higher boiling point and a lower freezing point

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**Arturo Cuomo**

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