

Predicting Products of Chemical Reactions

How can we predict the products of a chemical reaction? We can use the following guidelines to help us predict the products of a chemical reaction.

1. Write the chemical formula for each reactant and product. Use the following guidelines to help you write the chemical formula for each reactant and product: a. For ionic compounds, write the cation first and the anion second. b. For molecular compounds, write the element with the lower electronegativity first and the element with the higher electronegativity second. c. For acids, write the hydrogen atom first and the anion second.

2. Write the chemical formula for each product. Use the following guidelines to help you write the chemical formula for each product: a. For ionic compounds, write the cation first and the anion second. b. For molecular compounds, write the element with the lower electronegativity first and the element with the higher electronegativity second. c. For acids, write the hydrogen atom first and the anion second.

1. Synthesis	$A + B \rightarrow AB$	2. Decomposition	$AB \rightarrow A + B$	3. Single displacement	$A + BC \rightarrow AC + B$	4. Double displacement	$AB + CD \rightarrow AD + CB$
5. Combustion	$C_xH_y + O_2 \rightarrow CO_2 + H_2O$						
6. Acid-base reaction	$HA + BOH \rightarrow BA + H_2O$						
7. Redox reaction	$2H_2 + O_2 \rightarrow 2H_2O$						
8. Precipitation	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$						
9. Neutralization	$HCl + NaOH \rightarrow NaCl + H_2O$						
10. Oxidation-reduction	$2Fe + 3Cl_2 \rightarrow 2FeCl_3$						
11. Single displacement	$Zn + 2HCl \rightarrow ZnCl_2 + H_2$						
12. Double displacement	$CaCl_2 + Na_2CO_3 \rightarrow CaCO_3 + 2NaCl$						
13. Synthesis	$2H_2 + O_2 \rightarrow 2H_2O$						
14. Decomposition	$2H_2O \rightarrow 2H_2 + O_2$						
15. Single displacement	$Fe + 2HCl \rightarrow FeCl_2 + H_2$						
16. Double displacement	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$						
17. Redox reaction	$2Fe + 3Cl_2 \rightarrow 2FeCl_3$						
18. Precipitation	$CaCl_2 + Na_2CO_3 \rightarrow CaCO_3 + 2NaCl$						
19. Neutralization	$HCl + NaOH \rightarrow NaCl + H_2O$						
20. Oxidation-reduction	$2Fe + 3Cl_2 \rightarrow 2FeCl_3$						
21. Single displacement	$Zn + 2HCl \rightarrow ZnCl_2 + H_2$						
22. Double displacement	$CaCl_2 + Na_2CO_3 \rightarrow CaCO_3 + 2NaCl$						
23. Synthesis	$2H_2 + O_2 \rightarrow 2H_2O$						
24. Decomposition	$2H_2O \rightarrow 2H_2 + O_2$						
25. Single displacement	$Fe + 2HCl \rightarrow FeCl_2 + H_2$						
26. Double displacement	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$						

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Y Pai

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