

DRAWING COVALENT BONDS

Introduction:

A **covalent bond** forms between **nonmetals**. Covalent bonds involve **sharing pairs of electrons** between two atoms so that each atom gains a full outer valence shell. Sharing 1 pair of electrons creates a **single covalent bond**, sharing 2 pairs of electrons creates a **double covalent bond**, and sharing 3 pairs of electrons creates a **triple covalent bond**.



How to Draw Simple Covalent Bond Diagrams	
1.	Draw the electron dot diagrams for each element. <i>HINT: This will help you see the number of unpaired valence electrons for each element.</i>
2.	Count the total number of valence electrons between all atoms. <i>HINT: The total number of valence electrons should be 8 in most cases, except 2 hydrogens. If there are not 8, add more atoms of the element with the least number of valence electrons to start.</i>
3.	Determine which element is the central atom in the molecule, if needed. <i>HINT: The central atom is typically the atom with the largest valence number or the atom that there is only one of in the formula. If there are only 2 atoms in the formula, there is no central atom!</i>
4.	Pair up single electrons from different atoms to form covalent bonds and circle each pair. <i>HINT: Keep doing this until each single electron from all atoms have been paired up. You may need to use double or triple bonds to make this work. This step takes trial and error. If one arrangement is not working, try adding another element or trying a new structure. Practice makes perfect!</i>

Activity: Drawing Covalent Bond Diagrams:

- For each covalent compound, you have been given the element names and the chemical formula.
- Follow the steps listed above to create a covalent bonding diagram for the compound.
- Remember, this takes trial and error. The first 2 have been done for you as examples.

#	Nonmetal	Nonmetal	Formula	Bonding Diagram
A.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	chlorine $\cdot \ddot{\text{Cl}} \cdot$ <i>(must share 1)</i>	HCl <i>(Formula shows 1 H and 1 Cl)</i>	
B.	hydrogen $\text{H} \cdot$ <i>(must share 1)</i>	oxygen $\cdot \ddot{\text{O}} \cdot$ <i>(must share 2)</i>	H_2O <i>(Formula shows 2 H and 1 O)</i>	

Answer Key For Chemistry Covalent Bond Notes Ch 8

Arshad Iqbal



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Table of Contents Answer Key For Chemistry Covalent Bond Notes Ch 8

1. Understanding the eBook Answer Key For Chemistry Covalent Bond Notes Ch 8
 - The Rise of Digital Reading Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Chemistry Covalent Bond Notes Ch 8
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Personalized Recommendations
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 User Reviews and Ratings

- Answer Key For Chemistry Covalent Bond Notes Ch 8 and Bestseller Lists
- 5. Accessing Answer Key For Chemistry Covalent Bond Notes Ch 8 Free and Paid eBooks
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 Public Domain eBooks
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 eBook Subscription Services
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 Budget-Friendly Options
- 6. Navigating Answer Key For Chemistry Covalent Bond Notes Ch 8 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 Compatibility with Devices
 - Answer Key For Chemistry Covalent Bond Notes Ch 8 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Highlighting and Note-Taking Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Interactive Elements Answer Key For Chemistry Covalent Bond Notes Ch 8
- 8. Staying Engaged with Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Chemistry Covalent Bond Notes Ch 8
- 9. Balancing eBooks and Physical Books Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Chemistry Covalent Bond Notes Ch 8
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Setting Reading Goals Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Fact-Checking eBook Content of Answer Key For Chemistry Covalent Bond Notes Ch 8
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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