

CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

After reading (1) _____ solution _____ of solute is dissolved in solvent.

(2) _____ solute _____ is the (3) _____ solvent _____ and solute is the

(4) _____ solution. A solution that does not dissolve is
 (5) _____ insoluble. (6) _____ liquid _____ solution is the most common
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the
 two liquids are (7) _____ miscible. However, if one liquid is insoluble in another
 the liquids are (8) _____ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Use the diagram to answer the following questions.



1. Indicate in solute is when the solute is in a solution. What is solute?

Solute is the process of surrounding solute particles with solvent particles to form a solution.

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Chemistry: Matter and Change • Chapter 15

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CHAPTER 15 STUDY GUIDE FOR CONSTITUTANT MASTERY**Section 15.1** *Answers*

16. An sodium chloride dissolved in water. What happens is the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

17. Unlike the attraction of the water molecule around the sodium ion and chloride ion.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atom), because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ion and chloride ion? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solution process to occur.

19. Use these steps that the use of solvents may be increased.

Starting by shaking the solution, breaking the solute into smaller pieces, and heating the solvent.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute and solvent, write true or false.

True 16. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 17. Solubility is a property of the solute, amount of water the solute can dissolve in a given amount of solvent at a specified temperature and pressure.

False 18. Solvent molecules in liquid as the solute has a low heat of solution.

True 19. In a saturated solution, solute and crystallization are in equilibrium.

True 20. Additional solute can be dissolved in an unsaturated solution.

21. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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Aftab S. Hassan

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