

Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

Property	Acids	Bases
Conductivity in Aqueous Solution	very	very
Taste	Sour	Bitter
Feel	—	Slippery
Reaction with Metal	produces H_2 gas	—
Reaction with Carbonate	produces CO_2 gas	—
Colour on Litmus Paper	Red	Blue
Colour with Universal Indicator	Red	Purple
Colour with phenolphthalein	Colourless	pink

2. Indicate if each compound is an acid or base and name it.

	Acid/Base?	Name
HNO_3	<u>A</u>	<u>nitric acid</u>
$Mg(OH)_2$	<u>B</u>	<u>magnesium hydroxide</u>
H_3PO_4	<u>A</u>	<u>phosphoric acid</u>
H_2CO_3	<u>A</u>	<u>carbonic acid</u>
HF	<u>A</u>	<u>hydrofluoric acid</u>
$Fe(OH)_3$	<u>B</u>	<u>iron (iii) hydroxide</u>

3. Write the formula for each of the following acids/bases:

hydrochloric acid	<u>HCl</u>	hydrochloric acid	<u>HCl</u>
aluminum hydroxide	<u>$Al(OH)_3$</u>	sodium hydroxide	<u>NaOH</u>
lithium hydroxide	<u>LiOH</u>	calcium hydroxide	<u>$Ca(OH)_2$</u>
sulfuric acid	<u>H_2SO_4</u>	sulfurous acid	<u>H_2SO_3</u>

Answer Key Acids Bases

Edward P. Ortleb, Richard Cadice



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Table of Contents Answer Key Acids Bases

1. Understanding the eBook Answer Key Acids Bases
 - The Rise of Digital Reading Answer Key Acids Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Acids Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Acids Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Acids Bases
 - Personalized Recommendations
 - Answer Key Acids Bases User Reviews and Ratings
 - Answer Key Acids Bases and Bestseller Lists
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 - Highlighting and Note-Taking Answer Key Acids Bases
 - Interactive Elements Answer Key Acids Bases
- 8. Staying Engaged with Answer Key Acids Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Acids Bases
- 9. Balancing eBooks and Physical Books Answer Key Acids Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Acids Bases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Acids Bases
 - Setting Reading Goals Answer Key Acids Bases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Acids Bases
 - Fact-Checking eBook Content of Answer Key Acids Bases
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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