

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C_4H_{10}) in a limited amount of oxygen (O_2). If 67.2 L of CO_2 gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L } CO_2 \cdot \frac{(1 \text{ mol } CO_2)}{(22.4 \text{ L } CO_2)} = 3 \text{ mol } CO_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol } CO_2 \cdot \frac{(1.3 \text{ mol } O_2)}{(8 \text{ mol } CO_2)} = 4.875 \text{ mol } O_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol } O_2 \cdot \frac{(22.4 \text{ L } O_2)}{(1 \text{ mol } O_2)} = 109.2 \text{ L } O_2 = \boxed{109 \text{ L } O_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K_2SO_4 . What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g } Li \cdot \frac{(1 \text{ mol } Li)}{(7 \text{ g } Li)} = 2 \text{ mol } Li$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol } Li \cdot \frac{(1 \text{ mol } Li_2SO_4)}{(2 \text{ mol } Li)} = 1 \text{ mol } Li_2SO_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol } Li_2SO_4 \cdot \frac{(119 \text{ g } Li_2SO_4)}{1 \text{ mol } Li_2SO_4} = \boxed{119 \text{ Li}_2SO_4}$$

Answer Guide For Stoichiometry Calculations

Ying Liu

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Table of Contents Answer Guide For Stoichiometry Calculations

1. Understanding the eBook Answer Guide For Stoichiometry Calculations
 - The Rise of Digital Reading Answer Guide For Stoichiometry Calculations
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Guide For Stoichiometry Calculations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Guide For Stoichiometry Calculations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Guide For Stoichiometry Calculations
 - Personalized Recommendations
 - Answer Guide For Stoichiometry Calculations User Reviews and Ratings
 - Answer Guide For Stoichiometry Calculations and Bestseller Lists
5. Accessing Answer Guide For Stoichiometry Calculations Free and Paid eBooks

- Answer Guide For Stoichiometry Calculations Public Domain eBooks
- Answer Guide For Stoichiometry Calculations eBook Subscription Services
- Answer Guide For Stoichiometry Calculations Budget-Friendly Options

6. Navigating Answer Guide For Stoichiometry Calculations eBook Formats

- ePUB, PDF, MOBI, and More
- Answer Guide For Stoichiometry Calculations Compatibility with Devices
- Answer Guide For Stoichiometry Calculations Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Answer Guide For Stoichiometry Calculations
- Highlighting and Note-Taking Answer Guide For Stoichiometry Calculations
- Interactive Elements Answer Guide For Stoichiometry Calculations

8. Staying Engaged with Answer Guide For Stoichiometry Calculations

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answer Guide For Stoichiometry Calculations

9. Balancing eBooks and Physical Books Answer Guide For Stoichiometry Calculations

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answer Guide For Stoichiometry Calculations

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answer Guide For Stoichiometry Calculations

- Setting Reading Goals Answer Guide For Stoichiometry Calculations
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answer Guide For Stoichiometry Calculations

- Fact-Checking eBook Content of Answer Guide For Stoichiometry Calculations
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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