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web a write a balanced equation for the reaction $Al + O_2 \rightarrow Al_2O_3$ b determine the theoretical yield of Al_2O_3 160.0 g $Al: 1 \text{ mol} / 27 \text{ g/mol} = 0.037 \text{ mol}$ $Al_2O_3: 0.037 \text{ mol} \times 101.964 \text{ g/mol} = 3.8 \text{ g}$ c determine the percent yield yield actual $\times 100 = 260.0 \text{ g} / 3.8 \text{ g} \times 100 = 68.4\%$ theoretical 302.3 g

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