



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base in terms of the Arrhenius theory.
- Define strong and weak acids.

Definitions

- Hydrogen ion: H^+
- Hydrogen ion concentration: H^+ concentration
- Acid: a substance that dissociates in water to produce hydrogen ions
- Base: a substance that dissociates in water to produce hydroxide ions

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Definitions

- Strong acid: dissociates completely in water
- Weak acid: dissociates partially in water

Acids & Bases

Acids and bases are substances that dissociate in water to produce hydrogen ions and hydroxide ions, respectively.

Strong acids dissociate completely in water to produce hydrogen ions. The concentration of these ions is equal to the concentration of the acid.

Weak acids dissociate partially in water to produce hydrogen ions. The concentration of the acid is greater than the concentration of the hydrogen ions.

Strong bases dissociate completely in water to produce hydroxide ions. The concentration of these ions is equal to the concentration of the base.

Weak bases dissociate partially in water to produce hydroxide ions.

Strong Acids

- Hydrochloric acid
- Sulfuric acid
- Acetic acid
- Hydrobromic acid
- Hydroiodic acid
- Hydrofluoric acid
- Hydroperchloric acid
- Hydrochlorous acid
- Hydrobromous acid
- Hydroiodous acid

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