

112. The IUPAC name of



(a)

(A) 2,2-dimethyl-2,4-dimethylhexane
 (B) 2,2-dimethyl-4,4-dimethylhexane
 (C)



(D) 4,4-dimethyl-2,2-dimethylhexane
 (E) 2-methyl-4,4-dimethylhexane

Ans.
Sol.

The correct answer is 2,2-dimethyl-4,4-dimethylhexane.

Hence, (A) is correct.

Which of the following species exhibits the diamagnetic behaviour?

(A) NO_3^-
 (B) O_2
 (C) NO_2
 (D)

Ans.
Sol.

The correct option is NO_3^- .

This species has 18 e^- , which are filled in such a way that all molecular orbitals are fully filled, are diamagnetic.

NO_3^- : NO_2^- : O_2^- : NO_2 : NO_3 : O_2 : NO : O

Hence, (A) is correct.

The stability of anions of the XeO_4^- ion and its increasing stability in the sequence

(A) $\text{XeO}_4^- \rightarrow \text{BrO}_4^- \rightarrow \text{SbO}_4^- \rightarrow \text{PF}_6^-$
 (B) $\text{XeO}_4^- \rightarrow \text{ClO}_4^- \rightarrow \text{BF}_4^- \rightarrow \text{SO}_4^{2-}$
 (C) $\text{SO}_4^{2-} \rightarrow \text{ClO}_4^- \rightarrow \text{BrO}_4^- \rightarrow \text{XeO}_4^-$
 (D) $\text{PF}_6^- \rightarrow \text{BrO}_4^- \rightarrow \text{ClO}_4^- \rightarrow \text{SO}_4^{2-}$

Ans.
Sol.

Due to inert pair effect, the stability of $+2$ oxidation state increases as we move down the group.
 $\text{SO}_4^{2-} \rightarrow \text{ClO}_4^- \rightarrow \text{BrO}_4^- \rightarrow \text{XeO}_4^-$

Hence, (C) is correct.

Identify the incorrect statement among the following.

(A) Cl_2 reacts with NaOH to give NaOCl
 (B) S_8 reacts with NaOH in the presence of air to give $\text{Na}_2\text{S}_4\text{O}_6$ and H_2O
 (C) Cl_2 reacts with excess of NaOH to give NaCl and HClO
 (D) Br_2 reacts with hot and strong NaOH solution to give NaBrO_3 , NaBrO_4 and H_2O

Ans.
Sol.

Br_2 reacts with hot and strong NaOH to give NaBrO_3 , NaBrO_4 and H_2O .
 Hence, (D) is incorrect statement.

The charge/route ratio of a cation determines the polarising power. Which one of the following sequences represents the increasing order of the polarising power of the cationic species, K^+ , Ca^{2+} , Mg^{2+} , Be^{2+} ?

(A) Mg^{2+} , Be^{2+} , K^+ , Ca^{2+}
 (B) Be^{2+} , K^+ , Ca^{2+} , Mg^{2+}
 (C) K^+ , Ca^{2+} , Mg^{2+} , Be^{2+}
 (D) Ca^{2+} , Mg^{2+} , Be^{2+} , K^+

Ans.
Sol.

Higher the charge/route ratio, more is the polarising power.

$\text{K}^+ < \text{Ca}^{2+} < \text{Mg}^{2+} < \text{Be}^{2+}$

Hence, (D) is correct.

117. The density (in g mL^{-1}) of a 2.00 M sulphuric acid solution that is 20% H_2SO_4 (Molar mass = 98 g/mol) by mass will be

(A) 1.048
 (B) 1.032
 (C)

Ans.
Sol.

Let the density of solution be x g mL^{-1}

Molar mass of solution given = 98 g

i.e., 1 liter of solution contains 2.00 moles of H_2SO_4

or 1 liter of solution contains 2.0×98 grams of H_2SO_4

Since, this solution is 20% by mass.

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