112. The IUPAC name of

is

(1) 1, 1-diethyl-2,2-dimethylpentane   (2) 4, 4-dimethyl-5, 5-diethylpentane
(3) 5, 5-diethyl-4, 4-dimethylpentane   (4) 3-ethyl-4, 4-dimethylheptane

Ans. (4)
Sol.

The correct answer is 3-ethyl-4, 4-dimethylheptane.
Hence, (4) is correct.

113. Which of the following species exhibits the diamagnetic behaviour?
(1) O_2^+ (2) O_2^-
(3) O_2 (4) NO

Ans. (1)
Sol. The correct option is O_2^{2-}
This species has 18 e⁻, which are filled in such a way that all molecular orbitals are fully filled, so diamagnetic.
$\sigma 1s^2 \sigma^*1s^2 \sigma 2s^2 \sigma^*2s^2 \sigma 2p_z^2 \pi 2p_x^2 = \pi 2p_y^2 \pi^*2p_x^2 = \pi^*2p_y^2$
Hence, (1) is correct.

114. The stability of dihalides of Si, Ge, Sn and Pb increases steadily in the sequence
(1) $GeX_2 < SiX_2 < SnX_2 < PbX_2$ (2) $SiX_2 < GeX_2 < PbX_2 < SnX_2$
(3) $SiX_2 < GeX_2 < SnX_2 < PbX_2$ (4) $PbX_2 < SnX_2 < GeX_2 < SiX_2$

Ans. (3)
Sol. Due to inert pair effect, the stability of +2 oxidation state increases as we move down this group.
$\therefore SiX_2 < GeX_2 < SnX_2 < PbX_2$
Hence, (3) is correct.

115. Identify the incorrect statement among the following
(1) Ozone reacts with SO_2 to give SO_3.
(2) Silicon reacts with $NaOH_{(aq)}$ in the presence of air to give Na_2SiO_3 and H_2O.
(3) Cl_2 reacts with excess of NH_3 to give N_2 and HCl.
(4) Br_2 reacts with hot and strong NaOH solution to give NaBr, $NaBrO_3$ and H_2O.

Ans. (3)
Sol. Br_2 reacts with hot and strong NaOH to give NaBr, $NaBrO_3$ and H_2O.
Hence, (4) is incorrect statement.

116. The charge/size ratio of a cation determines its polarizing power. Which one of the following sequences represents the increasing order of the polarizing order of the polarizing power of the cationic species, $K^+, Ca^{2+}, Mg^{2+}, Be^{2+}$?
(1) $Mg^{2+} < Be^{2+} < K^+ < Ca^{2+}$ (2) $Be^{2+} < K^+ < Ca^{2+} < Mg^{2+}$
(3) $K^+ < Ca^{2+} < Mg^{2+} < Be^{2+}$ (4) $Ca^{2+} < Mg^{2+} < Be^{2+} < K^+$

Ans. (3)
Sol. Higher the charge/size ratio, more is the polarizing power.
$\therefore K^+ < Ca^{2+} < Mg^{2+} < Be^{2+}$
Hence, (3) is correct.

117. The density (in g mL⁻¹) of a 3.60 M sulphuric acid solution that is 29% H_2SO_4 (Molar mass = 98 g mol⁻¹) by mass will be:
(1) 1.64   (2) 1.88
(3) 1.22   (4) 1.45

Ans. (3)
Sol. Let the density of solution be 'd'
Molarity of solution given = 3.6
i.e. 1 litre of solution contains 3.6 moles of H_2SO_4
or 1 litre of solution contains 3.6 × 98 gms of H_2SO_4
Since, the solution is 29% by mass.

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