

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HBr~~ **hydrobromic acid**
- ~~HNO₃~~ **nitric acid**
- ~~H₂SO₄~~ **sulfuric acid**
- ~~HClO₄~~ **perchloric acid**
- ~~Sodium hydroxide~~ **NaOH**
- ~~potassium hydroxide~~ **KOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~barium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NaOH**

~~HClO₃ for chloric acid & HNO₂ for nitrous acid. Dissociate in water~~

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.

HCl

~~HBr~~

~~HNO₃~~

~~Mg(OH)₂~~

NaOH

NaOH

~~NaOH~~

NaOH

3. What is the formula for hydronium? **H₃O⁺**

4. Cu_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{Cu}_2\text{CO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3 + \text{Cu}_2\text{CO}_3$
- Would you expect a large or small equilibrium constant for this reaction? **small, because Cu_2CO_3 is a weak acid.**
- Would you define Cu_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is Cu_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of Cu_2CO_3 : **Cu_2CO_3**
- Identify Cu_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{F}^- + \text{H}_2\text{O} \rightarrow \text{HF} + \text{OH}^-$
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of F^- : **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HSO_4^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for this reaction: $\text{HSO}_4^- + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4 + \text{SO}_4^{2-}$
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO_4^- is a weak acid.**
- Would you define HSO_4^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HSO_4^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of HSO_4^- : **SO_4^{2-}**
- Identify HSO_4^- as polyprotic, amphoteric, or neither. **amphoteric, because it has a negative charge (and will like to accept a H^+) and it has H's (and will like to lose one).**

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AN Whitehead



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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations
 - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
 - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists

5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services
 - 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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