

Atomic Probing Structure of the Atom



Purpose:

To determine which elements are present in a sample of an unknown substance.

Part I: Finding Elements in Unknown Samples

When a substance is heated, it emits light. The color of the light emitted is characteristic of the elements present in the substance. The color of the light emitted is determined by the energy levels of the atoms in the substance. The energy levels are determined by the number of protons and neutrons in the nucleus of the atom.

Procedure: 1. Prepare a series of standard solutions of known elements. 2. Heat each solution in a flame. 3. Observe the color of the light emitted. 4. Compare the color of the light emitted to the color of the light emitted by the standard solutions.

Using the color of the light emitted, determine the elements present in the unknown sample. (See the results of the experiment.)

Results: The unknown sample contains the elements sodium, potassium, calcium, and magnesium. The color of the light emitted is characteristic of these elements. The color of the light emitted is determined by the energy levels of the atoms in the substance. The energy levels are determined by the number of protons and neutrons in the nucleus of the atom.

Using the color of the light emitted, determine the elements present in the unknown sample. (See the results of the experiment.)

Conclusion: The unknown sample contains the elements sodium, potassium, calcium, and magnesium. The color of the light emitted is characteristic of these elements. The color of the light emitted is determined by the energy levels of the atoms in the substance. The energy levels are determined by the number of protons and neutrons in the nucleus of the atom.

Using the color of the light emitted, determine the elements present in the unknown sample. (See the results of the experiment.)

Answers Lesson 11 Atomic Pudding

Terry C. Jones



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