

2000 AP® CHEMISTRY FREE-RESPONSE QUESTIONS

CHEMISTRY—SECTION II

(Total time—90 minutes)

Part A

Time—40 minutes

YOU MAY USE YOUR CALCULATOR FOR PART A.

CLEARLY SHOW THE METHOD USED AND STEPS INVOLVED IN ARRIVING AT YOUR ANSWERS. It is to your advantage to do this, because you may earn partial credit if you do and you will receive little or no credit if you do not. Attention should be paid to significant figures. Be sure to write all your answers to the questions on the lined pages following each question in this booklet.

Answer Question 1 below. The Section II score weighting for this question is 20 percent.



1. When heated, hydrogen sulfide gas decomposes according to the equation above. A 3.40 g sample of $\text{H}_2\text{S}(g)$ is introduced into an evacuated rigid 1.25 L container. The sealed container is heated to 483 K, and 3.72×10^{-2} mol of $\text{S}_2(g)$ is present at equilibrium.

(a) Write the expression for the equilibrium constant, K_c , for the decomposition reaction represented above.

(b) Calculate the equilibrium concentration, in mol L^{-1} , of the following gases in the container at 483 K.

- $\text{H}_2(g)$
- $\text{H}_2\text{S}(g)$

(c) Calculate the value of the equilibrium constant, K_c , for the decomposition reaction at 483 K.

(d) Calculate the partial pressure of $\text{S}_2(g)$ in the container at equilibrium at 483 K.

(e) For the reaction $\text{H}_2(g) + \frac{1}{2} \text{ S}_2(g) \rightleftharpoons \text{H}_2\text{S}(g)$ at 483 K, calculate the value of the equilibrium constant, K_c .

Answers Ap Chemistry 2013 Response Questions

Camilla Rothe

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Table of Contents Answers Ap Chemistry 2013 Response Questions

1. Understanding the eBook Answers Ap Chemistry 2013 Response Questions
 - The Rise of Digital Reading Answers Ap Chemistry 2013 Response Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Ap Chemistry 2013 Response Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Ap Chemistry 2013 Response Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Ap Chemistry 2013 Response Questions
 - Personalized Recommendations
 - Answers Ap Chemistry 2013 Response Questions User Reviews and Ratings
 - Answers Ap Chemistry 2013 Response Questions and Bestseller Lists
5. Accessing Answers Ap Chemistry 2013 Response Questions Free and Paid eBooks

- Answers Ap Chemistry 2013 Response Questions Public Domain eBooks
- Answers Ap Chemistry 2013 Response Questions eBook Subscription Services
- Answers Ap Chemistry 2013 Response Questions Budget-Friendly Options

6. Navigating Answers Ap Chemistry 2013 Response Questions eBook Formats

- ePUB, PDF, MOBI, and More
- Answers Ap Chemistry 2013 Response Questions Compatibility with Devices
- Answers Ap Chemistry 2013 Response Questions Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Answers Ap Chemistry 2013 Response Questions
- Highlighting and Note-Taking Answers Ap Chemistry 2013 Response Questions
- Interactive Elements Answers Ap Chemistry 2013 Response Questions

8. Staying Engaged with Answers Ap Chemistry 2013 Response Questions

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Answers Ap Chemistry 2013 Response Questions

9. Balancing eBooks and Physical Books Answers Ap Chemistry 2013 Response Questions

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Answers Ap Chemistry 2013 Response Questions

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Answers Ap Chemistry 2013 Response Questions

- Setting Reading Goals Answers Ap Chemistry 2013 Response Questions
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Answers Ap Chemistry 2013 Response Questions

- Fact-Checking eBook Content of Answers Ap Chemistry 2013 Response Questions
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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