

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
2. Calculate K_a and K_b for weak acids and bases, and calculate pK_a and pK_b for weak acids and bases.
3. Calculate the pH of weak acid and base solutions.

Vocabulary

- | | | |
|-----------------|--------------------------------------|------------------------------------|
| acid | acid dissociation constant (K_a) | acid ionization constant (K_a) |
| acidic solution | acidic solution | acidic solution |
| acidic solution | acidic solution | acidic solution |

Key Equations

- | | |
|---|--|
| $K_a = \frac{[H^+][A^-]}{[HA]}$ (Equation 19.1) | $K_b = \frac{[OH^-][HB^+]}{[B]}$ (Equation 19.2) |
|---|--|

Part A Completion

Write the K_a expression for the acid dissociation constant for each of the following acids. Write the K_b expression for the base dissociation constant for each of the following bases.

1. HNO_3 (strong acid) $K_a =$ _____

2. HCl (strong acid) $K_a =$ _____

3. H_2SO_4 (strong acid) $K_a =$ _____

4. H_2SO_3 (weak acid) $K_a =$ _____

5. H_2CO_3 (weak acid) $K_a =$ _____

6. $H_2PO_4^-$ (weak acid) $K_a =$ _____

7. $H_2PO_4^-$ (weak acid) $K_a =$ _____

8. $H_2PO_4^-$ (weak acid) $K_a =$ _____

9. $H_2PO_4^-$ (weak acid) $K_a =$ _____

10. $H_2PO_4^-$ (weak acid) $K_a =$ _____

11. $H_2PO_4^-$ (weak acid) $K_a =$ _____

12. HNO_3 (strong acid) $K_a =$ _____

13. HCl (strong acid) $K_a =$ _____

14. H_2SO_4 (strong acid) $K_a =$ _____

15. H_2SO_3 (weak acid) $K_a =$ _____

16. H_2CO_3 (weak acid) $K_a =$ _____

17. $H_2PO_4^-$ (weak acid) $K_a =$ _____

18. $H_2PO_4^-$ (weak acid) $K_a =$ _____

19. $H_2PO_4^-$ (weak acid) $K_a =$ _____

20. $H_2PO_4^-$ (weak acid) $K_a =$ _____

21. $H_2PO_4^-$ (weak acid) $K_a =$ _____

22. $H_2PO_4^-$ (weak acid) $K_a =$ _____

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