



Armed Forces Institute of Pathology

Advanced Laboratory Methods in Histology and Pathology

Ulrika V. Mikel, Editor

**Armed Forces Institute of Pathology
American Registry of Pathology
Washington, DC**

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R. Behme



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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{t}{t_{1/2}} \times \ln 2 \times \frac{50.3}{125 \text{ g}}$ what is the half life of an isotope that is 80 remained after 16 days solution remaining 80 100 therefore $n = \frac{t}{t_{1/2}} \times \ln 2 \times \frac{80}{100}$ now using the half life equation in example 1 we have

web what is the half life of this radioisotope after 2 days 5 grams of the radioisotope 1 day b 2 days c 5 days d 4 days 3 the ratio of uranium 238 to lead 206 in a mineral is used to determine age b density c solubility d composition 4 a radioactive isotope has a half life of 10 years

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