

**AP® CHEMISTRY
2013 SCORING GUIDELINES**

**Question 1
(10 points)**

Answer the following questions about the solubility of some fluoride salts of alkaline earth metals.

(a) A student prepares 100. mL of a saturated solution of MgF_2 by adding 0.50 g of solid MgF_2 to 100. mL of distilled water at 25°C and stirring until no more solid dissolves. (Assume that the volume of the undissolved MgF_2 is negligibly small.) The saturated solution is analyzed, and it is determined that $[\text{F}^-]$ in the solution is $2.4 \times 10^{-3} \text{ M}$.

(i) Write the chemical equation for the dissolving of solid MgF_2 in water.



(ii) Calculate the number of moles of MgF_2 that dissolved.

$\frac{2.4 \times 10^{-3} \text{ mol F}^-}{1.0 \text{ L}} \times 0.100 \text{ L} \times \frac{1 \text{ mol MgF}_2}{2 \text{ mol F}^-} = 1.2 \times 10^{-4} \text{ mol MgF}_2$	1 point is earned for the correct calculation of moles from concentration.
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(iii) Determine the value of the solubility-product constant, K_{sp} , for MgF_2 at 25°C .

$[\text{Mg}^{2+}] = \frac{1}{2} [\text{F}^-] = \frac{1}{2} (2.4 \times 10^{-3} \text{ M}) = 1.2 \times 10^{-3} \text{ M}$	1 point is earned for the correct value of $[\text{Mg}^{2+}]$.
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$$K_{sp} = [\text{Mg}^{2+}][\text{F}^-]^2 = (1.2 \times 10^{-3})(2.4 \times 10^{-3})^2$$
$$= 6.9 \times 10^{-9}$$

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