

Name \_\_\_\_\_

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

1) The conjugate base of  $\text{HSO}_4^-$  is \_\_\_\_\_.

A)  $\text{H}_2\text{SO}_4$       B)  $\text{SO}_4^{2-}$       C)  $\text{H}_2\text{SO}_4^+$       D)  $\text{HSO}_4^+$       E)  $\text{OH}^-$

2) The conjugate acid of  $\text{HSO}_4^-$  is \_\_\_\_\_.

A)  $\text{H}^+$       B)  $\text{H}_2\text{SO}_4$       C)  $\text{HSO}_4^+$       D)  $\text{SO}_4^{2-}$       E)  $\text{HSO}_3^-$

3) The molar concentration of hydroxium ions in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.

A) 7.00      B)  $1.0 \times 10^{-7}$       C) 1.00      D)  $1.0 \times 10^{-14}$       E) 0.00

4) The molar concentration of hydroxide ion in pure water at  $25^\circ\text{C}$  is \_\_\_\_\_.

A) 1.00      B)  $1.0 \times 10^{-14}$       C)  $1.0 \times 10^{-7}$       D) 0.00      E) 7.00

5) What is the pH of an aqueous solution at  $25^\circ\text{C}$  in which  $[\text{OH}^-]$  is 0.0025 M?

A) +2.60      B) -2.60      C) +11.40      D) -11.40      E) -2.25

6) What is the pH of an aqueous solution at  $25^\circ\text{C}$  that contains  $3.98 \times 10^{-9}$  M hydroxium ion?

A) 7.000      B) 9.000      C) 8.400      D) 5.600      E) 3.980

7) What is the concentration (in M) of hydroxium ions in a solution at  $25^\circ\text{C}$  with pH = 4.28?

A) 4.28      B)  $1.92 \times 10^{-10}$       C)  $1.66 \times 10^4$       D) 9.71      E)  $5.22 \times 10^{-5}$

8) Which solution below has the highest concentration of hydroxide ions?

A) pH = 3.21      B) pH = 9.82      C) pH = 7.93      D) pH = 1.259      E) pH = 7.00

9) What is the pH of a 0.015-M aqueous solution of barium hydroxide?

A) 12.18      B) 1.52      C) 12.48      D) 1.82      E) 10.35

10) Sodium hydroxide is a strong base. This means that \_\_\_\_\_.

A) aqueous solutions of NaOH contain equal concentrations of  $\text{H}^+$  (aq) and  $\text{OH}^-$  (aq)  
 B) NaOH does not dissociate at all when it is dissolved in water  
 C) NaOH dissociates completely to  $\text{Na}^+$  (aq) and  $\text{OH}^-$  (aq) when it dissolves in water  
 D) NaOH cannot be neutralized by a weak acid  
 E) NaOH cannot be neutralized by ordinary means

11) Of the following acids, \_\_\_\_\_ is not a strong acid.

A)  $\text{HNO}_3$       B)  $\text{HCl}$       C)  $\text{HNO}_2$       D)  $\text{HCO}_3$       E)  $\text{H}_2\text{SO}_4$

12) Of the following, \_\_\_\_\_ is a weak acid.

A)  $\text{HBr}$       B)  $\text{HF}$       C)  $\text{HClO}_4$       D)  $\text{HCl}$       E)  $\text{HNO}_3$

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**David Wilson**

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and Assessing Understanding Section VIII The book's final section IX is devoted to Professional Issues facing college and university faculty who choose to adopt active learning in their courses The common feature underlying all of the strategies described in this book is their emphasis on actively engaging students who seek to make sense of natural objects and events Many of the strategies we highlight emerge from a constructivist view of learning that has gained widespread acceptance in recent years In this view learners make sense of the world by forging connections between new ideas and those that are part of their existing knowledge base For most students that knowledge base is riddled with a host of naive notions misconceptions and alternative conceptions they have acquired throughout their lives To a considerable extent the job of the teacher is to coax out these ideas to help students understand how their ideas differ from the scientifically accepted view to assist as students restructure and reconcile their newly acquired knowledge and to provide opportunities for students to evaluate what they have learned and apply it in novel circumstances Clearly this prescription demands far more than most college and university scientists have been prepared for *5 Steps to a 5 AP Chemistry, 2012-2013 Edition* Richard H.

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## **Table of Contents 2013 Ap Chemistry Practice Test**

1. Understanding the eBook 2013 Ap Chemistry Practice Test
  - The Rise of Digital Reading 2013 Ap Chemistry Practice Test
  - Advantages of eBooks Over Traditional Books
2. Identifying 2013 Ap Chemistry Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an 2013 Ap Chemistry Practice Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from 2013 Ap Chemistry Practice Test
  - Personalized Recommendations
  - 2013 Ap Chemistry Practice Test User Reviews and Ratings
  - 2013 Ap Chemistry Practice Test and Bestseller Lists
5. Accessing 2013 Ap Chemistry Practice Test Free and Paid eBooks
  - 2013 Ap Chemistry Practice Test Public Domain eBooks
  - 2013 Ap Chemistry Practice Test eBook Subscription Services

- 2013 Ap Chemistry Practice Test Budget-Friendly Options
- 6. Navigating 2013 Ap Chemistry Practice Test eBook Formats
  - ePUB, PDF, MOBI, and More
  - 2013 Ap Chemistry Practice Test Compatibility with Devices
  - 2013 Ap Chemistry Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of 2013 Ap Chemistry Practice Test
  - Highlighting and Note-Taking 2013 Ap Chemistry Practice Test
  - Interactive Elements 2013 Ap Chemistry Practice Test
- 8. Staying Engaged with 2013 Ap Chemistry Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers 2013 Ap Chemistry Practice Test
- 9. Balancing eBooks and Physical Books 2013 Ap Chemistry Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection 2013 Ap Chemistry Practice Test
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine 2013 Ap Chemistry Practice Test
  - Setting Reading Goals 2013 Ap Chemistry Practice Test
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 2013 Ap Chemistry Practice Test
  - Fact-Checking eBook Content of 2013 Ap Chemistry Practice Test
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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