

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from  $\text{pH}$  and  $\text{pOH}$ , and vice versa from  $\text{pH}$  and  $\text{pOH}$ .
3. Calculate the  $\text{pH}$  of a weak acid.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. $\text{pH}$                | 6. $\text{pOH}$  |

## Key Equation

- |                                    |                                  |
|------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

## Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the weak acids and bases that are indicated in the spaces provided. Also indicate whether each acid or base is classified as weak.

What substance is \_\_\_\_\_, a weak hydrogen ion concentration \_\_\_\_\_.

Autoionization of \_\_\_\_\_ is the autoionization of \_\_\_\_\_.

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Autoionization

# 19 2 Hydrogen Ions And Acidity Work Sheet Answers

**Rod O'Connor**



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