

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or base solution.
2. Calculate pH values and concentrations of acids and bases and use these values to determine the pH of a given solution.
3. Calculate the pH of a buffer (see problems).

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. buffer solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the requested unit or numerical information of the values indicated that are indicated in the problem. Write the unit or numerical value in the space provided.

1. What substance is _____, a weak hydrogen ion donor?
2. _____ is the concentration of H^+ ions present in a solution and equal to _____.
3. _____ is the concentration of H^+ ions in a solution.
4. _____ is the concentration of H^+ ions in a solution.
5. _____ is the concentration of H^+ ions in a solution.
6. _____ is the concentration of H^+ ions in a solution.
7. _____ is the concentration of H^+ ions in a solution.
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17. _____ is the concentration of H^+ ions in a solution.
18. _____ is the concentration of H^+ ions in a solution.
19. _____ is the concentration of H^+ ions in a solution.
20. _____ is the concentration of H^+ ions in a solution.

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19 2 Hydrogen Ions And Acidity Answer Key

Arturo Cuomo



19 2 Hydrogen Ions And Acidity Answer Key:

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