

Chapter 1

Some Basic Concepts of Chemistry

Answer 1.9:

The average atomic mass of chlorine

$$\begin{aligned}
 &= [(\text{Fractional abundance of } {}^{35}\text{Cl}) (\text{Molar mass of } {}^{35}\text{Cl})] \\
 &+ [(\text{Fractional abundance of } {}^{37}\text{Cl}) (\text{Molar mass of } {}^{37}\text{Cl})] \\
 &= \left[\left\{ \left(\frac{75.77}{100} \right) (34.9689\text{u}) \right\} + \left\{ \left(\frac{24.23}{100} \right) (36.9659\text{u}) \right\} \right] \\
 &= 26.4959 + 8.9568 = 35.4527 \text{ u}
 \end{aligned}$$

The average atomic mass of chlorine = 35.4527 u

Question 1.10:

In three moles of ethane (C_2H_6), calculate the following:

1. Number of moles of carbon atoms.
2. Number of moles of hydrogen atoms.
3. Number of molecules of ethane.

Answer 1.10:

- 1 mole of C_2H_6 contains 2 moles of carbon atoms.
 \therefore Number of moles of carbon atoms in 3 moles of C_2H_6 = $2 \times 3 = 6$
- 1 mole of C_2H_6 contains 6 moles of hydrogen atoms.
 \therefore Number of moles of carbon atoms in 3 moles of C_2H_6 = $3 \times 6 = 18$
- 1 mole of C_2H_6 contains 6.023×10^{23} molecules of ethane.
 \therefore Number of molecules in 3 moles of C_2H_6
 $= 3 \times 6.023 \times 10^{23} = 18.069 \times 10^{23}$

Question 1.11:

What is the concentration of sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) in mol L⁻¹ if its 20g are dissolved in enough water to make a final volume up to 2L.

Answer 1.11:

Molarity (M) of a solution is given by,

$$\begin{aligned}
 &= \frac{\text{Number of moles of solute}}{\text{Volume of solution in Litres}} = \frac{\text{Mass of sugar}/\text{molar mass of sugar}}{2\text{L}} \\
 &= \frac{20\text{g}/[(12 \times 12) + (1 \times 22) + (11 \times 16)] \text{ g}}{2\text{L}} = \frac{20\text{g}/342 \text{ g}}{2\text{L}} = \frac{0.0585 \text{ mol}}{2\text{L}} = 0.02925 \text{ mol L}^{-1}
 \end{aligned}$$

\therefore Molar concentration of sugar = 0.02925 mol L⁻¹

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Sir Monier Monier-Williams

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