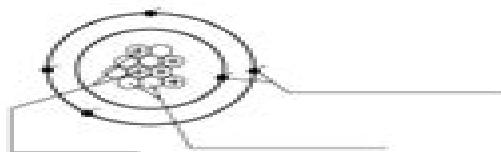


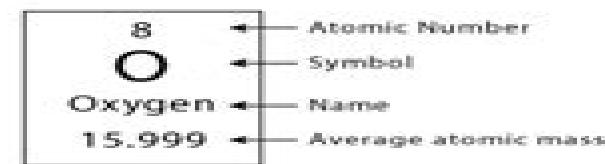
# Basic Atomic Structure

1. Label the parts of an atom in the diagram below



- What type of charge does a proton have? \_\_\_\_\_
- What type of charge does a neutron have? \_\_\_\_\_
- What type of charge does an electron have? \_\_\_\_\_
- Which two subatomic particles are located in the nucleus of an atom? \_\_\_\_\_

2. An element is represented by its chemical symbol along with a few numbers. Based on the example on the right, fill in the numbers of protons (P), neutrons (N), and electrons (E) for the following elements.



6 <b>C</b> Carbon 12.011	#P _____	#N _____	#E _____
-----------------------------------	----------	----------	----------

10 <b>Ne</b> Neon 20.108	#P _____	#N _____	#E _____
-----------------------------------	----------	----------	----------

19 <b>K</b> Potassium 39.098	#P _____	#N _____	#E _____
---------------------------------------	----------	----------	----------

3. Complete the table using your knowledge of the periodic table.

Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of Neutrons
Na			11		12
K		39		19	
			38		50
F				9	10
	20	40		20	
	50			50	69
I	53	127			

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**Emil J. Margolis**

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### **Table of Contents Atomic Structure Exercise 3 Answers**

1. Understanding the eBook Atomic Structure Exercise 3 Answers
  - The Rise of Digital Reading Atomic Structure Exercise 3 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Atomic Structure Exercise 3 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Atomic Structure Exercise 3 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Atomic Structure Exercise 3 Answers
  - Personalized Recommendations

- Atomic Structure Exercise 3 Answers User Reviews and Ratings
- Atomic Structure Exercise 3 Answers and Bestseller Lists

5. Accessing Atomic Structure Exercise 3 Answers Free and Paid eBooks

- Atomic Structure Exercise 3 Answers Public Domain eBooks
- Atomic Structure Exercise 3 Answers eBook Subscription Services
- Atomic Structure Exercise 3 Answers Budget-Friendly Options

6. Navigating Atomic Structure Exercise 3 Answers eBook Formats

- ePUB, PDF, MOBI, and More
- Atomic Structure Exercise 3 Answers Compatibility with Devices
- Atomic Structure Exercise 3 Answers Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Atomic Structure Exercise 3 Answers
- Highlighting and Note-Taking Atomic Structure Exercise 3 Answers
- Interactive Elements Atomic Structure Exercise 3 Answers

8. Staying Engaged with Atomic Structure Exercise 3 Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Atomic Structure Exercise 3 Answers

9. Balancing eBooks and Physical Books Atomic Structure Exercise 3 Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Atomic Structure Exercise 3 Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Atomic Structure Exercise 3 Answers

- Setting Reading Goals Atomic Structure Exercise 3 Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Atomic Structure Exercise 3 Answers

- Fact-Checking eBook Content of Atomic Structure Exercise 3 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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