

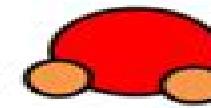
Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

1. The chemical formula shows:

- The _____ in the compound.
- The ratio of their _____.



B. Chemical Bond

1. Strong attractive force between _____ or ions in a molecule or compound.

2. Formed by:

- electrons, (e-⁻) by losing or gaining electrons.
- electrons (e-⁻)

C. Stability

1. Octet Rule

- Most atoms form bonds in order to have _____ valence electrons (e-⁻).
- outer energy level.
- Want to be like the Noble Gases (group 18).



2. Stability is the driving force behind bond formation. All atoms bond to become stable.

3. Transferring e-



4. Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

1. Attraction between 2 _____ charged ions.



2. Ions are _____ atoms.

3. A _____ is a _____ charged ion. Forms when an atom loses an electron.

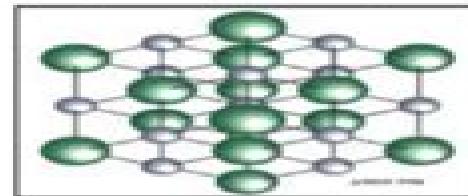
Formed by _____.

4. A _____ is a _____ charged ion. Forms when an atom gains an electron.

Formed by _____.

5. Formed by _____ electrons (e-⁻) from a _____ to a _____.

6. Ions form a 3-D crystal lattice.



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Archie L. Lacey, Virginia M. Rowley

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