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# 5100 White Corn Planter

**Rachel Sandford**

**5100 White Corn Planter:**

*Northern Wayne County Conservation Tillage Demonstration Project* Thomas DeRue, James Smith, 1986      **Union**  
**Agriculturist and Western Prairie Farmer**, 1999      *Wallaces Farmer*, 1999      Conservation Tillage for Corn Handbook  
University of Wisconsin--Extension. Cooperative Extension Service, 1986      *Evaluation Report*, 1983      **Northeast Region**  
**Official Guide**, 2001      *The Southern Planter*, 1910      **Farm Journal**, 2005      **Agri Finance**, 1990      *Journal of the*  
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California Farmer, 1993      **The American Herd Book** Lewis Falley Allen, 1875 To which is prefixed a concise history of  
English and American Short horns compiled from the best authorities      **Official Guide, Tractors and Farm Equipment**  
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**Horticulture and the Mining, Mechanic and Household Arts**, 1910      *Wallace's Farmer*, 2010      **Ohio Practical**  
**Farmer**, 1882

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element therefore differ from each other only in the number of neutrons within the nucleus  
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web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope's mass number

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