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R. H. Andrews

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web a reaction and its experimental rate law are represented below
 $2 \text{ ICl} \text{ (g)} \xrightarrow{\text{h}} 2 \text{ I} \text{ (g)} + 2 \text{ HCl} \text{ (g)}$ rate $k \text{ (s}^{-1}\text{)}$
a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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